

Equilibrium IN CLASS Practice Assessment

Skill: The student will be able to define equilibrium, draw and explain how a catalyst changes a reaction rate graph and discuss other changes that disrupt an equilibrium system

1. Complete the following:

a. What is "equal" at chemical equilibrium?

The rates of the forward & reverse reaction

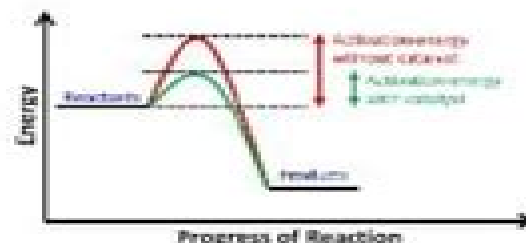
b. Provide one example of a factor that can disrupt equilibrium.

Concentration, temp, or pressure (if gases present)

c. Draw a new reaction pathway representing how a catalyst, like an enzyme, would affect the rate of a chemical reaction.

Provide an explanation for your new curve.

The addition of a catalyst lowers the activation energy necessary



Skill: The student will be able to write the equilibrium expression, given a reaction.

2. Write the equilibrium expression for the following reaction:



$$K = \frac{[\text{CaCl}_2][\text{CO}_2]}{[\text{HCl}]^2}$$

Skill: The student will be able to calculate the equilibrium constant and tell if the products or reactants are favored.

3. Calculate the equilibrium constant, for the reaction shown in Level 2, using the following equilibrium concentrations: $\text{CaCO}_3 = 1.0\text{M}$, $\text{HCl} = 0.20\text{M}$, $\text{H}_2\text{O} = 0.500\text{M}$, $\text{CaCl}_2 = 0.32\text{M}$ and $\text{CO}_2 = 0.42\text{M}$

$$K = \frac{[\text{CaCl}_2][\text{CO}_2]}{[\text{HCl}]^2} \rightarrow \frac{[.32][.42]}{[.20]^2} = 3.36$$

Which are favored, reactants or products? **Forward**

Skill: The student will be able to use Ksp correctly in an equilibrium calculation.

4. Calculate the Ksp, for iron (III) hydroxide, if the concentration is $4.49 \times 10^{-16}\text{M}$ at 25°C .



Skill: The student will be able to apply LeChatelier's principle correctly to a given situation and explain how the equilibrium will shift to compensate for this change, referring to collision theory.

5. To a sample of the cobalt equilibrium system studied in class, consider the following changes and say **what color** the system will turn? **Explain** using LeChatelier's principle & collision theory.



a. If magnesium chloride is dissolved in the system... **Right (Blue)**

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