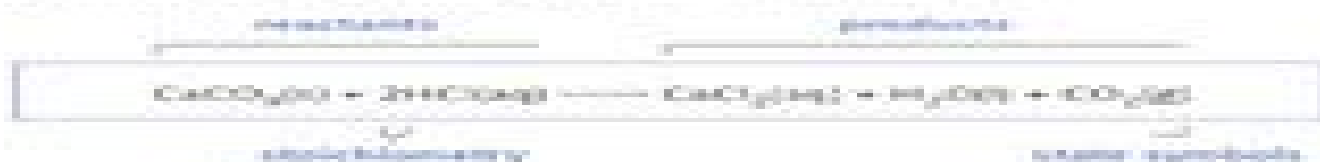


Writing Balanced Equations

Balanced chemical equations provide a lot of useful information about chemical reactions.

The reaction between solid calcium carbonate, CaCO_3 , and a solution of hydrochloric acid, HCl , produces a solution of calcium chloride, CaCl_2 , water, H_2O , and carbon dioxide gas, CO_2 . This reaction can be represented by the following equation:



This is known as a **full equation** because it shows the full chemical formula of all of the **reactants** (on the left-hand side) and all of the **products** (on the right-hand side).

The arrow in the equation represents the transformation that occurs during the chemical reaction.

The numbers in front of the formulae of the reactants and products in a balanced equation represent the **stoichiometry** of the reaction. This is the ratio in which reactants react with each other and products are produced. In the example above, one mole of calcium carbonate would react completely with two moles of hydrochloric acid to produce one mole of calcium chloride, one mole of water and one mole of carbon dioxide.

Finally, **state symbols** indicate the physical state of reactants and products during the chemical reaction: (s) represents a solid, (l) represents a liquid, (g) represents a gas, and (aq) represents an aqueous solution (a compound dissolved in water).

Balancing Equations

Atoms are not lost or gained in chemical reactions; they are simply rearranged to form different substances. Therefore, the total number of each type of atom represented on the left-hand side of an equation must equal the total number of each type of atom represented on the right-hand side of an equation. In the full equation above, the number of atoms of each element is the same on both sides of the equation, so the equation is balanced.

| Element | Number of Atoms of Element in Reactants | Number of Atoms of Element in Products |
|--------------|---|--|
| calcium, Ca | 1 | 1 |
| carbon, C | 1 | 1 |
| oxygen, O | 3 | 3 |
| hydrogen, H | 2 | 2 |
| chlorine, Cl | 2 | 2 |

If the equation is not balanced, you must change the stoichiometric ratios in the equation until it is balanced.

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