

# Concept Review

## Section 8-1: Covalent Bonds

Answer the following items in the space provided.

1. What does a molecular bond-forming reaction look like?
2. Why is the  $H_2$  molecule more stable than two separate hydrogen atoms?
3. Explain why the stability difference in  $H_2$  does not show that bond-forming is more energetically favorable.
4. What does a molecular bond-forming reaction look like?
5. What is enthalpy change?
6. Describe potential energy changes that occur when two hydrogen atoms approach each other, what forms a covalent bond?
7. In terms of energy, why is the  $H_2$  bond stable?
8. Why is bond length an average value? How is bond energy?
9. Describe the physical properties of molecules that the covalent bond, ionic, and covalent bonds.

Complete each statement below by choosing a term from the following list. Name may be used more than once.

- |                     |                     |          |                |
|---------------------|---------------------|----------|----------------|
| strategy            | electron-attraction | electron | ionic covalent |
| electron-attraction | electron            | electron | electron       |
1. Bond \_\_\_\_\_ is the force that average electron-attraction between two bonded atoms is. Bond \_\_\_\_\_ is the force that average electron-attraction between two bonded atoms is.
  2. Bond \_\_\_\_\_ is the force that average electron-attraction between two bonded atoms is. Bond \_\_\_\_\_ is the force that average electron-attraction between two bonded atoms is.
  3. The \_\_\_\_\_ of an atom is the average electron-attraction between two bonded atoms. The \_\_\_\_\_ of an atom is the average electron-attraction between two bonded atoms.
  4. In \_\_\_\_\_, electrons are attracted to the nucleus of the atom in which they are. In \_\_\_\_\_, electrons are attracted to the nucleus of the atom in which they are.
  5. In \_\_\_\_\_, electrons are attracted to the nucleus of the atom in which they are. In \_\_\_\_\_, electrons are attracted to the nucleus of the atom in which they are.

# Concept Review Chapter 6 Section Covalent Bonds

## Answer Key

**Holt Rinehart & Winston, Holt,  
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Organic Chemistry Michael B. Smith, 2010-10-18  
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George Plopper, 2014-10-21 Written for undergraduate cell biology courses Principles of Cell Biology Second Edition provides

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