

CHAPTER 19 REVIEW

# ***Oxidation-Reduction Reactions***

## **SECTION 2**

**SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ All of the following should be done in the process of balancing redox equations *except*:
  - (a) adjusting coefficients to balance atoms.
  - (b) adjusting coefficients in electron equations to balance numbers of electrons lost and gained.
  - (c) adjusting subscripts to balance atoms.
  - (d) writing two separate electron equations.
  
2.  $\text{MnCl}_2$  can be reduced to  $\text{MnO}_2$ .
  - \_\_\_\_\_ a. Assign the oxidation number to Mn in these two species.
  - \_\_\_\_\_ b. How many electrons are gained per Mn atom in this reduction?
  - \_\_\_\_\_ c. If 0.50 mol of  $\text{MnO}_2$  is reduced, how many electrons are gained?
  
3. Iodide ions can be oxidized to form iodine. Write the balanced oxidation half-reaction for the oxidation of iodide to iodine.
 

\_\_\_\_\_
  
4. Some bleaches contain aqueous chlorine as the active ingredient. Aqueous chlorine is made by dissolving chlorine gas in water. Aqueous chlorine is capable of oxidizing iron(II) ions to iron(III) ions. When iron(II) ions are oxidized, chloride ions are formed.
  - a. Write equations for the two half-reactions involved. Label them oxidation or reduction.
 

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\_\_\_\_\_
  - b. Write the balanced ionic equation for the redox reaction between aqueous chlorine and iron(II).
 

\_\_\_\_\_
  - c. Show that the equation in part b is balanced by charge.
 

\_\_\_\_\_

# Concept Review Section Oxidation Reduction Reactions Answers

**AN Whitehead**



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