

Section 6-1

Concept Review

Section: Simple Ions

Complete each statement below by choosing a term from the following list. Terms may be used more than once.

10

11

ion

octet

12

13

anion

cation

1. An atom or group of atoms that has a positive or negative electric charge because it has lost or gained electrons is a(n) ion.
2. The sodium ion has 11 protons, 10 neutrons, and 12 electrons.
3. In most chemical reactions, atoms tend to match the outer electron configuration of the noble gases. This is called the octet rule.

Complete each statement below by writing the correct term or phrase.

4. An anion is an ion with a negative charge.
5. Valence electrons occupy the outermost energy level of an atom.
6. The chloride ion and the chlorine atom have different electron configurations.
7. A cation is an ion with a positive charge.
8. Many stable ions have an electron configuration of a octet.

Answer the following questions in the space provided.

9. How do the outer-shell electron configurations for ions of Group 1, Group 2, and Group 15, Group 16, and Group 17 elements compare with those of the noble gases?

noble gases are stable while everything else is missing some

Concept Review Section Simple Ions Answers

William Mansfield Clark



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Principles of Life David M. Hillis,2012 For sample chapters a video interview with David Hillis and more information visit www.whfreeman.com/hillispreview Sinauer Associates and W H Freeman are proud to introduce Principles of Life Written in the spirit of the reform movement that is reinvigorating the introductory majors course Principles of Life cuts through the thicket of excessive detail and factual minutiae to focus on what matters most in the study of biology today Students explore the most essential biological ideas and information in the context of the field s defining experiments and are actively engaged in analyzing research data The result is a textbook that is hundreds of pages shorter and significantly less expensive than the current majors introductory books

Contemporary Chemical Analysis Judith F. Robinson,Kenneth A. Robinson,1998 This book provides a readable yet rigorous introduction to analytical methods with a focus on problem solving skills It stresses the fundamental concepts of chemical analysis and through examples from current journals and other science media shows how the principles and practice of analytical chemistry are used to produce answers to questions in all areas of scientific study and practice Features a balance of topics that is closer to contemporary analytical practice than those covered by other books Introduces the tools that are ubiquitous in analytical chemistry e g statistics sampling and sample preparation Discusses methods depending on chemical kinetics which are so widely used in medicine and biology Features a number of problems that call for the use of a spreadsheet to generate data which is then plotted to show trends Includes answers for all numerical problems in an appendix

Ion Exchange and Solvent Extraction Yitzhak Marcus,Arup K. SenGupta,2004-07-23 Volume 17 in the Ion Exchange and Solvent Extraction series represents the vanguard of research on solvent extraction It covers the principles of electrolyte extraction and other subjects of increasing interest to the field This volume begins with pharmaceutical applications of supercritical fluid solvents particularly supercritical carbon dioxide It also contains chapters on liquid ion

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