



...the rate of chemical reactions. The greater the concentration of reactant molecules, the more frequently they collide with one

this equation in words. How does this equation relate to breathing?

For suggested answers, see Appendix A.

2 Chapter Review

Go to **Mastering Biology** for Assignments, the eText, the Study Area, and Dynamic Study Modules.

SUMMARY OF KEY CONCEPTS

- To review key terms, go to the **Vocabulary Self-Quiz** in the **Mastering Biology** eText or Study Area, or go to go.jgk.org/2e.

CONCEPT 2.1

Matter consists of chemical elements in pure form and in combinations called compounds (pp. 29–30).

- Elements** cannot be broken down chemically to other substances. A **compound** contains two or more different elements in a fixed ratio. Oxygen, carbon, hydrogen, and nitrogen make up approximately 96% of living matter.

- Compare an element and a compound.

CONCEPT 2.2

An element's properties depend on the structure of its atoms (pp. 30–36).

- An **atom**, the smallest unit of an element, has the following components:



- An electrically neutral atom has equal numbers of electrons and protons; the number of protons determines the **atomic number**. The **atomic mass** is measured in **daltons** and is roughly equal to the **mass number**, the sum of protons plus neutrons. **Isotopes** of an element differ from each other in neutron number and therefore mass. Unstable isotopes give off particles and energy as **radioactivity**.
- In an atom, electrons occupy specific **electron shells**; the electrons in a shell have a characteristic energy level. Electron distribution in shells determines the chemical behavior of an atom. An atom that has an incomplete outer shell, the **valence shell**, is more reactive.
- Electrons exist in **orbitals**, three-dimensional spaces with specific shapes that are components of electron shells.

DISCUSS IT: Draw the electron distribution diagrams for neon (${}_{10}\text{Ne}$) and argon (${}_{18}\text{Ar}$). Use these diagrams to explain why these elements are chemically unreactive.

CONCEPT 2.3

The formation and function of molecules and ionic compounds depend on chemical bonding between atoms (pp. 36–40).

- Chemical bonds** form when atoms interact and complete their valence shells. **Covalent bonds** form when pairs of electrons are shared.



- Molecules** consist of two or more covalently bonded atoms. The attraction of an atom for the electrons of a covalent bond is its **electronegativity**. If both atoms are the same, they have the same electronegativity and share a **nonpolar covalent bond**. Electrons of a **polar covalent bond** are pulled closer to the more electronegative atom, such as the oxygen in H₂O.
- An **ion** forms when an atom or molecule gains or loses an electron and becomes charged. An **ionic bond** is the attraction between two oppositely charged ions.



- Weak interactions influence the shapes of large molecules, and help molecules adhere to each other. A **hydrogen bond** is an attraction between a hydrogen atom carrying a partial positive charge (δ^+) and an electronegative atom carrying a partial negative charge (δ^-). **Van der Waals interactions** occur between transiently positive and negative regions of molecules.
- A molecule's shape is determined by the positions of its atoms' valence orbitals. Covalent bonds result in hybrid orbitals, which are responsible for the shapes of H₂O, CH₄, and many more complex biological molecules. Molecular shape is usually the basis for the recognition of one biological molecule by another.

- In terms of electron sharing between atoms, compare nonpolar covalent bonds, polar covalent bonds, and the formation of ions.

CONCEPT 2.4

Chemical reactions make and break chemical bonds (pp. 40–41).

- Chemical reactions** change reactants into products while conserving matter. All chemical reactions are chemically reversible. **Chemical equilibrium** is reached when the forward and reverse reaction rates are equal.
- What would happen to the concentration of products if more reactants were added to a reaction that was in chemical equilibrium? How would this addition affect the equilibrium?

TEST YOUR UNDERSTANDING

- For more multiple-choice questions, go to the **Practice Test** in the **Mastering Biology** eText or Study Area, or go to go.jgk.org/2e.

Levels 1–2: Remembering/Understanding

- Compared with ${}^{12}\text{C}$, the radioactive isotope ${}^{14}\text{C}$ has
 - a different atomic number.
 - one more electron.
 - one more proton.
 - one more neutron.

Levels 5–6: Evaluating/Creating

- EVOLUTION CONNECTION** The percentages of naturally occurring elements making up the human body (see Table 2.1) are similar to the percentages of these elements found in other organisms. How could you account for this similarity among organisms?

- SCIENTIFIC INQUIRY** Pyruvate kinase mutase (PKM) is a protein that catalyzes the conversion of pyruvate to lactate. The enzyme organizes



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