

PROBLEM SET / CHEM 1411 — Ch 4: Chemical Reactions

Name _____

- 1) A 31.5 mL aliquot of H_2SO_4 (aq) of unknown concentration was titrated with 0.0134 M NaOH (aq). It took 23.9 mL of the base to reach the endpoint of the titration. What was the concentration (M) of the acid?
 - A) 0.227
 - B) 0.102
 - C) 0.0051
 - D) 0.0204
 - E) 0.0102
- 2) A stock solution of HNO_3 is prepared and found to contain 13.5 M of HNO_3 . If 25.0 mL of the stock solution is diluted to a final volume of 0.500 L, the concentration of the diluted solution is _____ M.
 - A) 675
 - B) 0.270
 - C) 0.675
 - D) 270
 - E) 1.48
- 3) What is the concentration of iodide ions in a 0.193 M solution of barium iodide?
 - A) 0.0965 M
 - B) 0.386 M
 - C) 0.193 M
 - D) 0.0643 M
 - E) 0.579 M
- 4) A strong electrolyte is one that _____ completely in solution.
 - A) disappears
 - B) dissolves
 - C) decomposes
 - D) reacts
 - E) ionizes
- 5) There are _____ mol of bromide ions in 0.500 L of a 0.300 M solution of AlBr_3 .
 - A) 0.150
 - B) 0.450
 - C) 0.167
 - D) 0.0500
 - E) 0.500
- 6) What is the concentration (M) of an aqueous methanol produced when 0.200 L of a 2.00 M solution was diluted to 0.800 L?
 - A) 8.00
 - B) 0.500
 - C) 0.800
 - D) 0.200
 - E) 0.400
- 7) Oxidation and _____ mean essentially the same thing.
 - A) activity
 - B) metathesis
 - C) corrosion
 - D) reduction
 - E) decomposition
- 8) When 0.500 mol of $\text{HC}_2\text{H}_3\text{O}_2$ is combined with enough water to make a 300 mL solution, the concentration of $\text{HC}_2\text{H}_3\text{O}_2$ is _____ M.
 - A) 3.33
 - B) 0.00167
 - C) 0.835
 - D) 1.67
 - E) 0.150
- 9) With which of the following will ammonium ion form an insoluble salt?
 - A) sulfate
 - B) chloride
 - C) carbonate
 - D) sulfate and carbonate
 - E) none of the above
- 10) When aqueous solutions of AgNO_3 and NaI are mixed, AgI precipitates. The balanced net ionic equation is _____
 - A) $\text{Ag}^+ (\text{aq}) + \text{I}^- (\text{aq}) \rightarrow \text{AgI} (\text{s})$
 - B) $\text{Ag}^+ (\text{aq}) + \text{NO}_3^- (\text{aq}) \rightarrow \text{AgNO}_3 (\text{s})$
 - C) $\text{Ag}^+ (\text{aq}) + \text{NO}_3^- (\text{aq}) \rightarrow \text{AgNO}_3 (\text{aq})$
 - D) $\text{AgNO}_3 (\text{aq}) + \text{NaI} (\text{aq}) \rightarrow \text{AgI} (\text{aq}) + \text{NaNO}_3 (\text{s})$
 - E) $\text{AgNO}_3 (\text{aq}) + \text{NaI} (\text{aq}) \rightarrow \text{AgI} (\text{s}) + \text{NaNO}_3 (\text{aq})$

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B Lingard



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