

Name Cierra Smith

Chemistry 151 Lab
Stoichiometric Relationships

Please work the following problems in the space provided and turn in at the end of the Lab period.

Mass Percent

1. Calculate the mass percent of Cl in Freon-12, a refrigerant with the formula $C_2Cl_4F_2$.

$$\begin{array}{r} 2C \times 12.01 = 24.02 \\ 4Cl \times 35.45 = 141.80 \\ 2F \times 19 = 38 \\ \hline C_2Cl_4F_2 = 203.82 \text{ g/mol} \end{array} \quad \frac{141.80}{203.82} \times 100 = \boxed{69.57\%}$$

2. The U.S. Food and Drug Administration recommends that you consume less than 2.4g of sodium per day. What mass of sodium chloride (in grams) can you consume and still be within the FDA guidelines? First determine the % by mass of sodium in sodium chloride. atomic mass of Na = 22.99

$$\begin{array}{l} \text{molar mass of NaCl} = 22.99 + 35.5 = 58.49 \\ \% \text{ by mass of Na} = \frac{22.99}{58.49} \times 100 = 39.30\% \end{array}$$

$$\frac{2.4 \times 58.49}{39} = \boxed{6.1 \text{ g}}$$

Empirical and Molecular Formulas

3. A compound containing nitrogen and oxygen decomposed in the laboratory and produces 24.5 g nitrogen and 70.0 g oxygen. Calculate the empirical formula of the compound.

$$\begin{array}{l} \text{mass of } N_2 = 24.5 \text{ g} \\ \text{molar mass of } N_2 = 28.01 \text{ g/mol} \\ \text{mass of } O_2 = 70 \text{ g} \\ \text{molar mass} = 32 \text{ g/mol} \end{array} \quad \begin{array}{l} \# \text{ of moles} \\ 24.5 \text{ g } N_2 \times \frac{1 \text{ mol } N_2}{28.01 \text{ g}} = 0.875 \\ 70 \text{ g } O_2 \times \frac{1 \text{ mol}}{32 \text{ g/mol}} = 2.1875 \end{array}$$

$$\begin{array}{l} N_2 \quad 0.875 \times 2 = 2 \\ O_2 = \frac{2.1875}{0.875} = 2.5 \times 2 = 5 \end{array} \quad \text{Empirical Formula: } \boxed{N_2O_5}$$

4. A laboratory analysis of aspirin determined the following mass percent composition: Carbon 60.00%; Hydrogen 4.48%; Oxygen 35.52%. Find its empirical formula.

$$\begin{array}{l} C = \frac{60}{12.01} = 4.996 / 2.22 = 2.25 \times 4 = 9 \\ H = \frac{4.48}{1.008} = 4.44 / 2.22 = 2 \times 4 = 8 \\ O = \frac{35.52}{16.00} = 2.22 / 2.22 = 1 \times 4 = 4 \end{array}$$

Empirical formula
 $C_9H_8O_4$

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