

## Observation Table

| S.No | Burette readings |       | net vol <sup>n</sup> used |
|------|------------------|-------|---------------------------|
|      | Initial          | Final |                           |
| (1)  | 0.0              | 14.4  | 14.4                      |
| (2)  | 14.4             | 29    | 14.6                      |
| (3)  | 29               | 43.6  | 14.6                      |

### Calculation

$$N_1 V_1 = N_2 V_2$$

(acid) (alk)

$$\frac{0.1 \times 10}{10} = N_2$$

$$N_2 = \frac{0.1 \times 14.6}{10} = 0.146 \text{ g/L}$$

$$\text{Strength} = N_{\text{NaOH}} \rightarrow 5.84 \text{ g/L}$$

total g/L

$$\text{strength of NaOH} = (8 - 1) = 7$$

$$8 - 5.84 = 2.16 \text{ g/L}$$

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### Experiment no.1

Aim: To find out the percentage composition of the given mixture of NaOH and NaCl.

Apparatus required: Burette, pipette, chemical flask, beaker.

Chemicals required: 0.1 N NaCl, NaOH, NaOH, NaOH, Phenolphthalein, indicator.

End point: Pink to colourless.

Reaction involved:



### Procedure

1) Rinse the bottle with the given NaOH sol<sup>n</sup> and standard found immediately.

2) Take 10 ml sol<sup>n</sup> in the burette and record the initial readings.

3) Pipette out the 10 ml of the solution in the chemical flask.

4) Add a drop of indicator in the solution.

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