

1. A solution is made by dissolving some salt in a beaker of water. The salt is referred to as the
A. solute B. filtrate C. solution D. solvent
2. When KCl dissolves in water, the following will be produced.
A. K and Cl B. K^+ and Cl^- C. K and Cl_2 D. K^+ and Cl_2
3. What is the concentration, in % (m/v) of a solution containing 15.0 g KCl in 600.0 mL solution?
A. 5.00% B. 2.00% C. 0.200% D. 2.50%
4. A 5.00 mL sample of solution has 2.8×10^{-4} g of calcium ions. The ppm concentration is
A. 18 ppm B. 56 ppm C. 2.8×10^{-5} ppm D. 2.8×10^7 ppm
5. How many mL of 2.50% (m/v) salt solution would contain 1.30 g of salt?
A. 105 mL B. 104 mL C. 45.0 mL D. 72.0 mL
6. What is the molarity of a solution that contains 1.50 mol HCl in 2.50 L of solution?
A. 1.67 M B. 0.600 M C. 1.20 M D. 1.40 M
7. How many moles of $C_{10}H_{22}O_{11}$ are needed to prepare 2.50 L of 0.300 M solution?
A. 0.750 mol B. 0.430 mol C. 8.33 mol D. 1.20 mol
8. What is the molarity of a $NaNO_3$ solution made by diluting 250.0 mL of a 1.60 M solution to a final volume of 400. mL?
A. 1.20 M B. 1.00 M C. 0.200 M D. 0.160 M
9. In the process known as osmosis, _____ moves through a semi-permeable membrane into an area of lower _____ concentration.
A. water, water B. solute, solute C. water, solute D. solute, water
10. What is the osmolarity of 0.800M $(NH_4)_3PO_4$ solution?
A. 3.20 osmol B. 2.40 osmol C. 1.60 osmol D. 8.00 osmol

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Henry Kiddle, Alexander Jacob Schem



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Chemical Pedagogy Keith S Taber, 2024-12-20 How should chemistry be taught in schools colleges and universities Chemical Pedagogy discusses teaching approaches and techniques the reasoning behind them and the evidence for their effectiveness The book surveys a wide range of different pedagogic strategies and tactics that have been recommended to better engage learners and provide more effective chemistry teaching These accounts are supported by an initial introduction to some key ideas and debates about pedagogy the science of teaching Chemical Pedagogy discusses how teaching innovations can be tested to inform research based practice Through this book the author explores the challenges of carrying out valid experimental studies in education and the impediments to generalising study results to diverse teaching and learning contexts As a result the author highlights both the need to read published studies critically and the value of teachers and lecturers testing out recommended innovations in their own classrooms Chemical Pedagogy introduces core principles from research into human cognition and learning to provide a theoretical perspective on how to best teach for engagement and understanding An examination of some of the more contentious debates about pedagogy leads to the advice to seek optimally guided instruction which balances the challenge offered to learners with the level of support provided This provides a framework for discussing a wide range of teaching approaches and techniques that have been recommended to those teaching chemistry across educational levels including both those intended to replace teaching from the front and others that can be built into traditional lecture courses to enhance the learning experience

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