

Honors Chemistry I
Electron Arrangement/Quantum Numbers Worksheet

1. What is the frequency of light (EM radiation) if it is known to have a wavelength of 832 **cm**? (Frequency is in units of /sec). [Hint: make sure to use the proper units!]

$$832 \text{ cm} = 8.32 \text{ m} \quad 3.00 \times 10^8 = 8.32(\nu) \quad \nu = 3.61 \times 10^7 \text{ s}^{-1}$$

- What amount of energy is associated with a photon at this wavelength?

$$E = 6.636 \times 10^{-34} (3.61 \times 10^7) \quad E = 2.40 \times 10^{-26} \text{ J}$$

- What amount of energy is associated with a mole of photons at this wavelength?

$$2.40 \times 10^{-26} \text{ J/photon} \times (6.022 \times 10^{23} \text{ photon/1 mole}) = 1.44 \times 10^{-2} \text{ J/mol}$$

2. What element is it whose neutral, isolated atom has two valence electrons in the 5s subshell and four electrons in the 5p subshell?

tellurium

3. How many energy levels (**shells**) are **either filled or partially filled** in the ground state configuration for the neodymium atom [Nd: Atomic #60]?

6

4. Based on our knowledge of the energy levels, what would be the predicted atomic number for the next noble gas (VIII-A) after radon (Neglect the hypothetical "g" subshell)?

118

5. How many **valence** electrons are there in an arsenic atom?

5

6. Which of the following is an electron configuration of an atom in the excited state? (May be more than 1)

A) $1s^2 2s^1$; B) $1s^2 2s^2 2p^1$; **C) $1s^2 2s^1 3d^1$;** **D) $1s^2 2s^1 2p^2$;** **E) [Ar] $3d^1$.**

7. How many sublevels are either filled or partially filled in an atom of iodine in the ground state?

$1s, 2s, 2p, 3s, 3p, 3d, 4s, 4p, 4d, 5s, 5p$

11

8. What is the total number of electrons in an atom if the **fourth shell** has been completely filled? Ignore any abnormalities in electron configuration, but remember the order as indicated by the Diagonal Rule and Aufbau Principle.

70

9. Which quantum number value indicates the most probable distance of the electron from the nucleus of an atom?

n

10. The **f** sublevel contains how many electrons? **14** How many orbitals? **7**

11. What is the next subshell to be filled immediately after **6s**?

4f

12. How many filled 3p orbitals are there in a neutral atom of cobalt?

3

13. What type of subshell is being filled to make the transition metals?

d

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