

Balancing Chemical Equations

(What stays the same and what may change in a chemical reaction?)

Model 1: Atoms are conserved in chemical reactions

Atoms are neither created nor destroyed when chemical reactions occur. Two balanced chemical reactions (or equations) are given below. In a chemical reaction equation, reactant atoms are shown on the left side of the arrow and product atoms are shown on the right.



Molecular Picture:



Critical Thinking Questions:

- Considering Equation (1) above:
 - What are the reactants? Write your answer in both words and formulas. Check that your teammates agree.
 - What is the product? Write your team's answer in both words and formulas.
- What does the arrow represent in a chemical reaction?
- How many hydrogen molecules are represented on the reactant side of equation (1)? ____
- How many water molecules are represented on the product side of equation (1)? ____
- How many H atoms are represented on the reactant side of equation (1)? ____
On the product side? ____
- How many O atoms are represented on the reactant side of equation (1)? ____
On the product side? ____
- Compare the number of atoms of each element (H and O) on the reactant and product sides of equation (1). Are these numbers **the same** or **different** (circle one)?
- Compare the number of atoms of each element (Fe, Al, and O) on the reactant and product sides of equation (2). Are these numbers **the same** or **different** (circle one)?
- In reaction (1):
 - For each H_2 molecule consumed, how many H_2O molecules are produced? ____
 - For each O_2 molecule consumed, how many H_2O molecules are produced? ____
 - Recall that a mole is simply a specific large quantity, like a dozen or a gross. For each mole of oxygen molecules consumed, how many moles of water molecules are produced? ____

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