

KEY

NAME \_\_\_\_\_ PER \_\_\_\_\_

**Lewis Structures and naming practice**

Naming covalent compounds is easy IF you remember six simple prefixes and one suffix. All we do to name covalently bound molecules is tell how many of the different atoms there are in the molecule by adding a prefix, then the whole thing ends with the *-ide* suffix. For example,  $\text{SiO}_2$  is silicon *dioxide*.  $\text{N}_2\text{O}_4$  is *dinitrogen tetroxide*.  $\text{CCl}_4$  is carbon *tetrachloride*. The six you need to know are –

**Di** – to indicate there are two of them

**Tri** – for three

**Tetra** – for four (sometimes we drop the “a” for tetra through deca (10), almost always when the element starts with a vowel– you’ll get a feel for it, and I’ll never mark you down if you don’t)

**Penta** – for five

**Hexa** – for six

**Mono** – for one (only used if there’s one of some element and multiples of another.

Then, you put the mono in for the singleton. For example,  $\text{CCl}_3\text{Br}$  is carbon trichloro, monobromide.)

I promise naming’s way easier than it seems. Try it with these as you practice Lewis structures. Write the name **or** the formula, and draw the Lewis structure.

<p>Dihydrogen oxide <math>2e^-</math></p> <p><math>\text{H}-\ddot{\text{O}}-\text{H}</math></p> <p><i>You should've known</i></p>	<p>Sulfur trioxide <math>24e^-</math></p> <p><math>\begin{array}{c} \ddot{\text{O}} \\ \parallel \\ \text{:}\ddot{\text{O}}-\text{S}-\ddot{\text{O}}\text{:} \end{array}</math></p>
<p><math>\text{SO}_2</math> Sulfur Dioxide <math>18e^-</math></p> <p><math>\text{:}\ddot{\text{O}}=\ddot{\text{S}}-\ddot{\text{O}}\text{:}</math></p> <p>★</p>	<p><math>\text{NH}_2\text{Cl}</math> <math>5+2(1)+7=14e^-</math> Nitrogen dichydrogen chloride</p> <p><math>\begin{array}{c} \text{H}-\ddot{\text{N}}-\text{H} \\   \\ \text{:}\ddot{\text{Cl}}\text{:} \end{array}</math></p>
<p><math>\text{NO}</math> Nitrogen oxide <math>11e^-?</math></p> <p><math>\text{:}\ddot{\text{N}}=\ddot{\text{O}}\text{:}</math></p> <p>★</p>	<p><math>\text{C}_2\text{F}_4</math> <math>2(4)+4(7)=36e^-</math></p> <p><math>\begin{array}{c} \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \\ \diagdown \quad \diagup \\ \text{:}\ddot{\text{F}}-\text{C}=\text{C}-\text{F}\text{:} \\ \diagup \quad \diagdown \end{array}</math></p> <p>Dicarbon tetrofluoride</p>

# Chemfiesta Lewis Structures Answer Key

**SJ Ball**



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## **Table of Contents Chemfiesta Lewis Structures Answer Key**

1. Understanding the eBook Chemfiesta Lewis Structures Answer Key
  - The Rise of Digital Reading Chemfiesta Lewis Structures Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemfiesta Lewis Structures Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemfiesta Lewis Structures Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemfiesta Lewis Structures Answer Key
  - Personalized Recommendations
  - Chemfiesta Lewis Structures Answer Key User Reviews and Ratings
  - Chemfiesta Lewis Structures Answer Key and Bestseller Lists
5. Accessing Chemfiesta Lewis Structures Answer Key Free and Paid eBooks

- Chemfiesta Lewis Structures Answer Key Public Domain eBooks
- Chemfiesta Lewis Structures Answer Key eBook Subscription Services
- Chemfiesta Lewis Structures Answer Key Budget-Friendly Options
- 6. Navigating Chemfiesta Lewis Structures Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemfiesta Lewis Structures Answer Key Compatibility with Devices
  - Chemfiesta Lewis Structures Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemfiesta Lewis Structures Answer Key
  - Highlighting and Note-Taking Chemfiesta Lewis Structures Answer Key
  - Interactive Elements Chemfiesta Lewis Structures Answer Key
- 8. Staying Engaged with Chemfiesta Lewis Structures Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemfiesta Lewis Structures Answer Key
- 9. Balancing eBooks and Physical Books Chemfiesta Lewis Structures Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemfiesta Lewis Structures Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemfiesta Lewis Structures Answer Key
  - Setting Reading Goals Chemfiesta Lewis Structures Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemfiesta Lewis Structures Answer Key
  - Fact-Checking eBook Content of Chemfiesta Lewis Structures Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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