

CHM 116 LAB PROCEDURE

Chemical Equilibrium: Calculating K_a of a Weak Acid

(~1 – 2 hours)

Introduction

In this experiment, you will determine the equilibrium constant, K_a , of a weak, monoprotic acid. The technique used in this experiment is that of **titration**, which is the slow addition of a reactant of known concentration (the **titrant**) to another reactant (the **analyte**) with the purpose of determining an unknown quantity concerning the analyte. The titration you will be performing is one where the titrant is a 0.1 M **sodium hydroxide (NaOH)** solution that will be titrated against **acetic acid (CH₃COOH)** while the pH of the reaction mixture is monitored throughout the procedure. The **endpoint** of the reaction (the point at which an excess of the **titrant (NaOH)** causes an indicator to change color) will be determined through the observation of a turmeric indicator. The reaction that is taking place in the reaction vessel is:



You will be determining the K_a (equilibrium constant of a weak acid) and the $\text{p}K_a$ for the acetic acid using pH and concentration data. Additionally, you will be constructing a **titration curve** for the reaction of a **weak acid (CH₃COOH)** with a **strong base (NaOH)**, by plotting pH (on the y-axis) against the volume of NaOH added (on the x-axis). The titration curve allows us to determine the equivalence point of the reaction (the point at which the moles of **CH₃COOH** present and the moles of **NaOH** added are equal). The **equivalence point** is the point on the curve which is half-way up the steep (closest to vertical) portion of the curve. The endpoint and the equivalence point are NOT the same thing, though they should be very close in value.

*Note: The weak acid, **acetic acid**, that you will be using in this lab is an organic acid, and as such the formula is often not written with the same convention as inorganic acids, where the acidic hydrogen is the first element in the formula. In organic acids, the formula is generally written with the acidic hydrogen at the end of the formula, after an oxygen. Thus, the formula for acetic acid is conventionally written as:*



*Where the final **H** is the one which will ionize in water.*

Chemical Equilibrium Lab Report Answers

Puja Mehta



Chemical Equilibrium Lab Report Answers:

Chemical Equilibrium Lab Report Answers Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Chemical Equilibrium Lab Report Answers**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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