

CHEMICAL EQUILIBRIUM UNIT REVIEW
(Chemistry 12)

I. Multiple Choice (60 marks)

1. 6. Consider the following reversible reaction: $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$

A solution of $\text{Fe}(\text{NO}_3)_3(\text{aq})$ is added to a solution of KSCN. Which one of the following statements describes the changes in forward and reverse reaction rates as the reaction moves towards equilibrium?

- A. Forward and reverse rates increase.
B. Forward and reverse rates decrease.
C. Forward rate increases and reverse rate decreases.
D. Forward rate decreases and reverse rate increases.

2. A system at equilibrium is said to be dynamic because at equilibrium the:

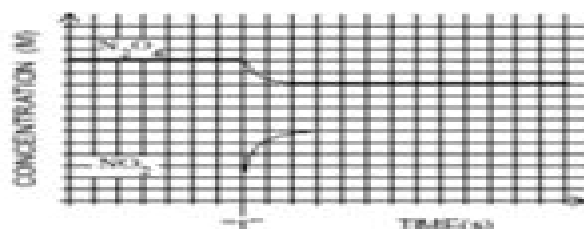
- A. temperature does not change.
B. macroscopic properties are constant.
C. forward and reverse reactions continue to occur.
D. concentrations of reactants and products are constant

3. Consider the following equilibrium system: $\text{CaCO}_3(\text{s}) \rightleftharpoons \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$

Which one of the following changes would cause the above system to shift left?

- A. Add more CaO
B. Remove CaCO_3 .
C. Decrease volume
D. Increase surface area of CaO

4. Consider the above concentration versus time graph for the equilibrium:



At time = " t ", which one of the following stresses occurred?

- A. Catalyst was added.
B. Pressure was changed.
C. Temperature was changed.
D. Concentration of NO_2 was changed.

Chemical Equilibrium Review Questions Answers

Multiple

John T. Moore, Richard H. Langley



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