

Chapter 18

Study Guide - Answers

1. Write the expression for the equilibrium constant for each of the following reactions.



$$K_{eq} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$



$$K_{eq} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$



$$K_{eq} = \frac{[\text{H}^+][\text{CN}^-]}{[\text{HCN}]}$$



$$K_{eq} = \frac{[\text{H}^+][\text{HSO}_4^-]}{[\text{H}_2\text{SO}_4]}$$



$$K_{eq} = \frac{[\text{Pb}^{2+}][\text{F}^-]^2}{1}$$

2. From the data provided below, calculate the value of the equilibrium constant for the reaction.



$$[\text{H}_2] = [\text{Cl}_2] = 1.0 \times 10^{-2}; [\text{HCl}] = 1.0 \times 10^{-4}$$

$$K_{eq} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$

$$K_{eq} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-2}][1.0 \times 10^{-2}]} = 0.0001$$



$$[\text{N}_2] = 4.4 \times 10^{-2}; [\text{H}_2] = 1.2 \times 10^{-1}; [\text{NH}_3] = 3.4 \times 10^{-3}$$

$$K_{eq} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$

$$K_{eq} = \frac{[3.4 \times 10^{-3}]^2}{[4.4 \times 10^{-2}][1.2 \times 10^{-1}]^3} = 0.152$$

3. For the reaction $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$, $K_{eq} = 0.2$. At a particular time, the following concentrations are measured: $[\text{N}_2\text{O}_4] = 2.0 \text{ M}$, $[\text{NO}_2] = 0.2 \text{ M}$. Is this reaction at equilibrium? If not, in which direction will the reaction proceed?



$$[\text{N}_2\text{O}_4] = 2.0 \text{ M}$$

$$[\text{NO}_2] = 0.2 \text{ M}$$

$$Q = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

$$Q = \frac{[0.2]^2}{[2.0]} = 0.02$$

$$K_{eq} = 0.2$$

$Q < K_{eq}$, so the reaction will proceed to the right (forward).

Chemical Equilibrium Study Guide Answers

Paul B. Kelter, Steven S. Zumdahl



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