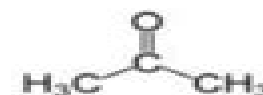


Unit 0 – Chemistry Review

Section 3 – Chemical Formulas



Molecular (Covalent) Compounds

- What information can we tell about a compound given its empirical formula?
We can tell the smallest whole-number ratio of all the atoms in a given compound.
- What additional information is conveyed when we know its molecular formula?
We can tell the mass of the molecule.
- What additional information do we know when we see its structural formula?
We can tell what each atom is bonded to.
- Two compounds have the same empirical formula; one is a gas at room temperature and the other is a liquid. How is this possible? Just because they have the same empirical formula does not mean they are the same compound. Small changes in atom arrangement and quantity can result in large differences in physical properties.
- Determine the molecular and empirical formula of the following: Molecular first, then empirical.
 - The organic solvent benzene that has six carbon atoms and six hydrogen atoms C_6H_6 CH
 - The compound silicon tetrachloride, a substance used to make computer chips SiCl_4 SiCl_4
 - The reactive substance diborane, a compound with two boron atoms and six hydrogen atoms B_2H_6 BH_3
 - The sugar glucose, a substance with six carbon, twelve hydrogen, and six oxygen atoms $\text{C}_6\text{H}_{12}\text{O}_6$ CH_2O
- Write down the molecular and structural formulas represented by the following models:



Notice how some are written slightly different. They are all correct representations even though one is more condensed and one shows the valence electrons.

Ionic Compounds

- Which of the following compounds below likely represents an ionic compound and which represents a covalent compound? Explain your answer.

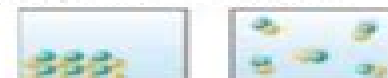
The first one is likely an ionic compound because of the repeating pattern (of positive and negative ions).

- Predict the charge of the most stable ion each element below is capable of forming:

a. Mg $2+$
b. Al $3+$

c. S $2-$
d. K $1+$

e. F $1-$
f. O $2-$



g. N $3-$
h. Na $1+$

- Fill in the table for the compound made from each cation and anion pair. Use parentheses when necessary.

	Cl^-	NO_3^-	S^{2-}	CO_3^{2-}	N^{3-}	PO_4^{3-}	OH^-
Na^+	NaCl	NaNO_3	Na_2S	Na_2CO_3	Na_3N	Na_3PO_4	NaOH
NH_4^+	NH_4Cl	NH_4NO_3	$(\text{NH}_4)_2\text{S}$	$(\text{NH}_4)_2\text{CO}_3$	$(\text{NH}_4)_3\text{N}$	$(\text{NH}_4)_3\text{PO}_4$	NH_4OH
Sn^{2+}	SnCl_2	$\text{Sn}(\text{NO}_3)_2$	SnS	SnCO_3	Sn_3N_2	$\text{Sn}_3(\text{PO}_4)_2$	$\text{Sn}(\text{OH})_2$
Hg_2^{2+}	Hg_2Cl_2	$\text{Hg}_2(\text{NO}_3)_2$	Hg_2S	Hg_2CO_3	$(\text{Hg}_2)_3\text{N}_2$	$(\text{Hg}_2)_3(\text{PO}_4)_2$	$\text{Hg}_2(\text{OH})_2$
Al^{3+}	AlCl_3	$\text{Al}(\text{NO}_3)_3$	Al_2S_3	$\text{Al}_2(\text{CO}_3)_3$	AlN	AlPO_4	$\text{Al}(\text{OH})_3$
Sn^{4+}	SnCl_4	$\text{Sn}(\text{NO}_3)_4$	SnS_2	$\text{Sn}(\text{CO}_3)_2$	Sn_3N_4	$\text{Sn}_3(\text{PO}_4)_4$	$\text{Sn}(\text{OH})_4$

NOTE: ONE of the above formulas is incorrect. If you can spot it, send me an e-mail. The first one to find it and email me the answer will receive extra credit!

Chemical Formulas And Compounds Chapter Review

Section 1

Iris Stovall, Steven S. Zumdahl



Chemical Formulas And Compounds Chapter Review Section 1:

Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science Study Guide for Chemical Principles Thomas Elliott Taylor, 1979 Organometallic Chemistry Ian J S Fairlamb, Jason M Lynam, 2011-06-07 A series of critical reviews and perspectives focussing on specific aspects of organometallic chemistry interfacing with other fields of study are provided For this volume the critical reviews cover topics such as the activation of inert carbon hydrogen bonds ligand design and organometallic radical species For example Charlie O Hara discusses how mixed metal compounds may perform the highly selective activation of C H bonds and in particular how synergic relationships between various metals are crucial to this approach The chemistry of a remarkable series of air stable chiral primary phosphine ligands is discussed in some depth by Rachel Hiney Arne Ficks Helge M3ller Bunz Declan Gilheany and Lee Higham This article focuses on the preparation of these ligands and also how they may be applied in various catalytic applications Bas De Bruin reports on how ligand radical reactivity can be employed in synthetic organometallic chemistry and catalysis to achieve selectivity in radical type transformations As well as highlighting ligand centered radical transformations in open shell transition metals an overview of the catalytic mechanism of Co II catalysed olefin cyclopropanation is given showing that enzyme like cooperative metal ligand radical reactivity is no longer limited to real enzymes Valuable and informative comprehensive reviews in the field of organometallic chemistry are also covered in this volume For example organolithium and organocuprate chemistry are reviewed by Joanna Haywood and Andrew Wheatley aspects in Group 2 Be Ba and Group 12 Zn Hg compounds by Robert Less Rebecca Melen and Dominic Wright metal clusters by Mark Humphrey and Marie Cifuentes and recent developments in the chemistry of the elements of Group 14 focusing on low coordination number compounds by Richard Layfield This volume therefore covers many synthetic and applied aspects of modern organometallic chemistry which ought to be of interest to inorganic organic and applied catalysis fields *Hazmat Chemistry Study Guide (Second Edition)* Jill Meryl Levy, 2011

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