

# Pre-Lab:

## Experiment 10; Chemical Kinetics

1. In an iodination of acetone reaction the following solutions make up the reaction mixture:

5 mL 4.0 M acetone + 5 mL 0.0050 M  $I_2$  + 5 mL 1.0 M HCl + 10 mL  $H_2O$

- a. What were the initial concentrations of acetone,  $I_2$  and  $H^+$  ion in the reaction mixture?

[acetone] 0.0008 M;  $[I_2]$   $1 \times 10^{-6}$  M;  $[H^+]$  0.0002 M

- b. It took 221 seconds for the  $I_2$  color to disappear from the reaction mixture. What was the rate of the reaction?

$$R = \frac{[I_2]}{t} = \frac{1 \times 10^{-6}}{221 \text{ sec}} = 4.5 \times 10^{-9}$$

- c. Write the rate equation as it would apply to this reaction mixture:

$$\text{Rate} = k[1 \times 10^{-6}]^a[0.0008]^b[0.0002]^c$$

# Chemical Kinetics Pre Lab Answers

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**Exploring General Chemistry in the Laboratory** Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

**International Journal of Chemical Kinetics**, 1987 Topics covered in this publication include quantitative relationships between molecular structure and chemical activity organic inorganic chemistry biochemical kinetics and reaction mechanisms Surface kinetics are also explored

**Chemistry in the Laboratory** James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

**University of Nebraska-Lincoln, Catalog: ARTS & SCIENCES, COLLEGE OF**. University of Nebraska--Lincoln. College Of Arts & Sciences, 1928

**Book of Abstracts** American Chemical Society. Meeting, American Chemical Society, 1974

**Integrated Approach to Coordination Chemistry** Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07 Coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules This book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry It not only highlights the key fundamental components of the coordination chemistry field it also exemplifies the historical development of concepts in the field In order to graduate as a chemistry major that fills the requirements of the American Chemical Society a student needs to take a laboratory course in inorganic chemistry Most professors who teach and inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine the environment molecular biology organic synthesis and inorganic materials

**Using Physical Models of Biomolecules to Teach Concepts of Biochemical Structure in Introductory Undergraduate Chemistry** John Yi, 2004

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of Technology. College of Engineering and Science,1956  
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University,1970      **Illinois Chemistry Teacher** ,1992  
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