

ii)  $\text{Mg}^{2+}$

iii)  $\text{N}^{3-}$

**Q-8:** What information did Kossel provide about chemical bonding?

**Q-9:** State with reasons

a) Which is more acidic - anhydrous  $\text{HCl}$  or aqueous  $\text{HCl}$ ?

b) Which is more polar -  $\text{CO}_2$  or  $\text{N}_2\text{O}$ ?

**Q-10:**  $\text{AlCl}_3$  is covalent when anhydrous. Determine whether it would remain covalent or change to an ionic state in an aqueous solution using the information provided below. Ionisation energy for  $\text{AlCl}_3 = 5137 \text{ kJ/mol}$ ,  $\Delta H_{\text{hydration}}$  for  $\text{Al}^{3+} = -4665 \text{ kJ/mol}$ ,  $\Delta H_{\text{hydration}}$  for  $\text{Cl}^- = -381 \text{ kJ/mol}$ .

**Q-11:** Draw the molecular structures of  $\text{OsF}_6$  and  $\text{XeF}_4$  using the VSEPR theory, highlighting the location of lone pairs of electrons and the hybridisation of the central atom.

**Q-12:** Determine the  $\text{HCl}$ 's ionic character percentage. Given that the  $\text{HCl}$ 's bond length is  $1.275$  and the observed dipole moment is  $1.83 \text{ D}$ .

**Q-13:** Describe the change in hybridisation, if any.

a) Of  $\text{Al}$  atom in the reaction  $\text{AlCl}_3 + \text{Cl}^- \rightarrow \text{AlCl}_4^-$ ?

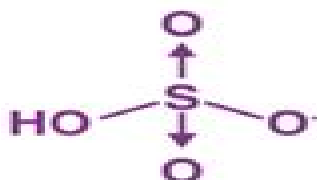
b) Of  $\text{B}$  and  $\text{N}$  atoms in the reaction  $\text{BF}_3 + \text{NH}_3 \rightarrow \text{F}_3\text{B}-\text{NH}_3$ ?

**Q-14:** Describe how Lewis' concept differs from VB theory.

**Q-15:** Out of  $sp$ -hybrid orbital and  $p$  orbital, which has greater directional character and why?

**Q-16:** When we state that the lattice energy of  $\text{NaCl}$  is  $788 \text{ kJ/mol}$ . What does it mean?

**Q-17:** Calculate the formal charge of  $\text{S}$  in the following structure.



**Q-18:** Give two resonating structures of ozone which satisfy the octet rule. Also, give the probable hybrid structure.

**Q-19:** i) Why does  $\text{MgCl}_2$  have a much higher solubility than  $\text{MgF}_2$ ?

ii) Why is the melting point of  $\text{MgO}$  ( $2800^\circ\text{C}$ ) much higher than that of  $\text{BaO}$  ( $1920^\circ\text{C}$ )?

**Q-20:** How will you differentiate between the lone pair and a bond pair of electrons?

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