CHAPTER 7 REVIEW

Chemical Formulas and Chemical Compounds

SECTION 7-1

SHORT ANSWER Answer the following questions in the space provided.

- In a Stock name such as iron(III) sulfate, the roman numeral tells us ______
 - (a) how many atoms of Fe are in one formula unit
 - (b) how many sulfate ions can be attached to the iron atom.
 - (c) the charge on each Fe ion
 - (d) the total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to ______
 - (a) change the number of moles represented by the formula
 - (b) change the charges on the other ions in the compound
 - (c) change the formula so that it no longer represents the compound it previously represented
 - (d) have no effect on the formula
- The explosive TNT has the molecular formula C₂H₃(NO₂)₃.
 - 4 elements a, How many elements make up this compound?
 - 6 oxygen atoms b. How many oxygen atoms are present in one molecule of $C_2H_3(NO_2)_3$?
 - 21 atoms c. How many atoms in total are present in one molecule of C₇H₅(NO₂)₃?
 - 4.2 × 10²⁴ atoms d. How many atoms are present in a sample of 2 × 10²³ molecules of C₇H₂(NO₂)₃?
- 4. How many atoms are present in each of these formula units?
 - 11 atoms a, Ca(HCO₃)₂
 - 45 atoms b. C₁₂H₂₂O₁₁
 - 10 atoms c. Fe(ClO₂)₃
 - 9 atoms d. Fe(ClO₃)₂
- 5. N₂Q₅ a. What is the formula for the compound dinitrogen pentoxide?
 - carbon(IV) sulfide b. What is the Stock name for the covalent compound CS₂?
 - H₂SO₃ c. What is the formula for sulfurous acid?
 - phosphoric acid d. What is the name for the acid H₃PO₄?

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Clifford Lane

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to treating the various categories of chemical reactions The organization is designed to get students into chemistry quickly yet methodically The classification of matter begins in Chapter 2 before the chapter on measurement and the coverage of bonding appears in Chapter 8 The elementary concepts of chemistry are presented with an emphasis on mathematical calculations and problem solving strategies Essentials of Chemistry Ralph A. Burns, John William Hill, 1995 The second edition of this chemistry textbook that uses practice examples and applications relating chemistry to our lives and the Basic Chemistry G. William Daub, William S. Seese, 1996 Written in a style and language that users without environment science backgrounds can understand This best selling introduction to the basic principles of chemistry draws on the reader s own experiences through analogies and cartoons to learn difficult concepts The clear systematic thinking approach to problem solving has also been highly praised by reviewers and users alike Countdown sections in each chapter consisting of five review questions keyed to previous material provide readers with a basis for material introduced in the new chapter Study exercises found immediately after new topics are introduced reinforce chapter problem material You and Chemistry marginal application icon relates chemistry to the real world End of chapter essays entitled Elements and Compounds relate the applications of specific elements or compounds to the readers life Robinson Chemistry Study Guide Robinson, 1992

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Table of Contents Chemical Names And Formulas Chapter 9 Practice Problems

- 1. Understanding the eBook Chemical Names And Formulas Chapter 9 Practice Problems
 - The Rise of Digital Reading Chemical Names And Formulas Chapter 9 Practice Problems
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemical Names And Formulas Chapter 9 Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Names And Formulas Chapter 9 Practice Problems
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemical Names And Formulas Chapter 9 Practice Problems

- Personalized Recommendations
- Chemical Names And Formulas Chapter 9 Practice Problems User Reviews and Ratings
- Chemical Names And Formulas Chapter 9 Practice Problems and Bestseller Lists
- 5. Accessing Chemical Names And Formulas Chapter 9 Practice Problems Free and Paid eBooks
 - Chemical Names And Formulas Chapter 9 Practice Problems Public Domain eBooks
 - Chemical Names And Formulas Chapter 9 Practice Problems eBook Subscription Services
 - Chemical Names And Formulas Chapter 9 Practice Problems Budget-Friendly Options
- 6. Navigating Chemical Names And Formulas Chapter 9 Practice Problems eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chemical Names And Formulas Chapter 9 Practice Problems Compatibility with Devices
 - Chemical Names And Formulas Chapter 9 Practice Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Names And Formulas Chapter 9 Practice Problems
 - Highlighting and Note-Taking Chemical Names And Formulas Chapter 9 Practice Problems
 - Interactive Elements Chemical Names And Formulas Chapter 9 Practice Problems
- 8. Staying Engaged with Chemical Names And Formulas Chapter 9 Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Names And Formulas Chapter 9 Practice Problems
- 9. Balancing eBooks and Physical Books Chemical Names And Formulas Chapter 9 Practice Problems
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Chemical Names And Formulas Chapter 9 Practice Problems
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Names And Formulas Chapter 9 Practice Problems
 - Setting Reading Goals Chemical Names And Formulas Chapter 9 Practice Problems
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Names And Formulas Chapter 9 Practice Problems

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- Fact-Checking eBook Content of Chemical Names And Formulas Chapter 9 Practice Problems
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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