

Chemical Names and Formulas

❖ Key Vocabulary

Monoatomic Ion – Ions formed from a single atom

Binary compound – Compounds composed of two elements

Nomenclature – the “naming system” of binary ionic compounds

Oxyanion – polyatomic ions that contain oxygen

Salt – an ionic compound composed of a cation and the anion from an acid

Chemical Formulas

Subscripts after an elemental symbol represent the number of atoms of the element in the compound

Molecular formulas represent how many atoms of each element are in a compound.
(C_2H_6 , H_2O , $C_6H_{12}O_6$)

Ionic formulas show the simplest ratio, cations to anions, of atoms in the lattice

Parenthesis are used around a polyatomic ion in a chemical formula

Examples $Al_2(SO_4)_3$, $Cu(PO_4)_2$, $Sr(NO_3)_2$

Monoatomic Ions

– Ions formed from a single atom

Group 1 elements may lose one electron = 1+

Group 2 elements may lose two electrons = 2+

Group 13 elements may lose three electrons = 3+

Group 17 elements may gain one electron = 1-

Group 16 elements may gain two electrons = 2-

Group 15 elements may gain three electrons = 3-

Transition metals have various possibilities (no pattern)

Naming monatomic Ions

Monatomic cations are simply called by the element's name

*the stock system is used when there is more than one possibility for an element (more on this later)

Monatomic anions end in IDE such as Fluorine – Fluoride, Oxygen – Oxide
(reference table 1 on page 221)

Binary Ionic Compounds

1. Write the symbols for the ions with the cation first (Al^{3+} O^{2-})

2. Cross over the charges writing the charge of one ion as the subscript for the other ion (Al_2^{3+} O_3^{2-})

3. Multiply subscript and superscript for each ion then find the lowest whole number ratio (Al_2O_3)

Name it!!! Aluminum Oxide

Examples: Zn^{2+} I^-

Zn^{2+} S^{2-}

(More on p 223 practice problems)

Chemical Names And Formulas Study Guide

James E. Brady, John R. Holum



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A Pronouncing Chemical Formula Speller and Contest Guide Carl Alfred Jacobson, 1925 **Study Guide for Chang's Chemistry** Kenneth W. Watkins, 1988 **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement **Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney** William H. Myers, William T. Mooney, 1980
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Introductory Chemistry Study Guide Holder, 2003-05 *General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material **Robinson Chemistry Study Guide** Robinson, 1992 **Science Information News**, 1961 **Study Guide for Introductory Chemistry** Iris Stovall, 1990 *General Chemistry, Study Guide* James E. Brady, 1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems **O Level Chemistry Questions and Answers PDF** Arshad Iqbal, The O Level Chemistry Quiz Questions and Answers PDF IGCSE GCSE Chemistry Competitive Exam Questions Chapter 1 14 Practice Tests Class 9 10 Chemistry Textbook Questions for Beginners includes revision guide for problem solving with hundreds of solved questions O Level Chemistry Questions and Answers PDF book covers basic concepts analytical and practical assessment tests O Level Chemistry Quiz PDF book helps to practice test questions from exam prep notes The O Level Chemistry Quiz Questions and Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved tests O Level Chemistry Questions and Answers PDF

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