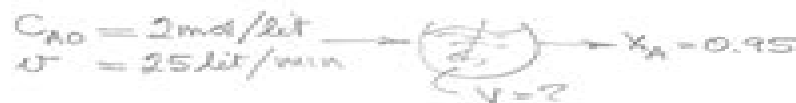


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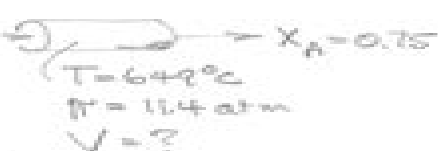
$$-r_A = \frac{0.1 C_A}{1 + 0.5 C_A} \text{ mol/lit} \cdot \text{min}$$

First of all $C_A = C_{A0}(1 - X_A) = 2(1 - 0.95) = 0.1$

Then for mixed flow

$$V = \frac{v(C_{A0} - C_A)}{-r_A} = \frac{25(2 - 0.1)}{\frac{0.1(0.1)}{1 + 0.5(0.1)}} = 4987.5 \text{ lit} \approx 5 \text{ m}^3 \leftarrow$$

5.13



$$-r_{\text{plus}} = (10 \text{ hr}^{-1}) C_{\text{plus}}$$

For a 1st order reaction

$$k \frac{C_{A0} V}{F_{A0}} = (1 + E_A) \ln \frac{1}{1 - X_A} - E_A X_A \quad \text{--- (i)}$$

Evaluate terms

$$C_{A0} = \frac{P_{A0}}{RT} = \frac{11.4 \times \frac{2}{3}}{(0.08206)(649 + 273)} = 0.1 \frac{\text{mol}}{\text{lit}}$$

$$k = 10 \text{ hr}^{-1}$$

$$F_{A0} = 10 \text{ mol/hr}$$

$$E_A = 0.5$$



$$\therefore E_A = \frac{90 - 60}{60} = \frac{1}{2}$$

Replacing in (i) gives

$$V = \frac{10}{10(0.1)} \left[(1 + 0.5) \ln \frac{1}{0.25} - 0.5(0.75) \right] = 17 \text{ lit} \leftarrow$$

5.15



$$\text{Evaluate terms } E_A = \frac{1-2}{2} = -0.5$$

$$X_A = \frac{C_{A0} - C_A}{C_{A0} + E_A C_A} = \frac{1 - 0.5}{1 + (-0.5)(0.5)} = \frac{2}{3}$$

So for mixed flow

$$v = \frac{V(-r_A)}{C_{A0} X_A} = \frac{(2)(0.05 \times 0.5^2)}{1(2/3)} = 0.0375 \text{ lit/s} = 2.25 \text{ lit/min} \leftarrow$$

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Chi M. Phan



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