

Chemical Reactions and Equations

10

A. Magnesium and Oxygen

1. Initial appearance of Mg Shiny light gray ribbon of metal
2. Evidence of chemical reaction Heat and light energy produced
White ash produced
3. Balance: $2 \text{Mg}_{(s)} + \text{O}_{2(g)} \xrightarrow{\Delta} 2 \text{MgO}_{(s)}$
4. Type of chemical reaction: Combination reaction

B. Zinc and Copper(II) Sulfate

Time	$\text{CuSO}_4(aq)$ Appearance	$\text{Zn}(s)$ Appearance	Evidence of a Chemical Reaction
1. initial	Blue/green	Dull gray	
2. after 30 min	Lt blue	Black	Change in color

3. $\text{Zn}_{(s)} + \text{CuSO}_{4(aq)} \xrightarrow{\Delta} \text{Cu}_{(s)} + \text{ZnSO}_{4(aq)}$ (the equation is balanced)

4. Type of chemical reaction: Single replacement reaction

C. Reactions of Metals and HCl

Metal	1. Appearance of Metals	2. Evidence of a Chemical Reaction
$\text{Zn}(s)$	Gray, shiny pebbles	Formation of bubbling
$\text{Cu}(s)$	Orange, shiny strips	No change
$\text{Mg}(s)$	Light gray, shiny strips	Formation of many bubbles Heat produced

Chemical Reactions And Equations Laboratory Manual

William Woolsey Johnson



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