

Name: _____

Tear this page out of your laboratory manual and submitted to your TA with your answers

B. Post-Lab Assessment Questions: (30 pts)

1. **(4 pts)** Write out the chemical reaction that occurs when an aqueous solution of iron(III) nitrate reacts with an aqueous solution of potassium iodide to produce an aqueous solution of iron(III)iodide and an aqueous solution of potassium nitrate.
2. **(3 pts)** Is the reaction in question 1 an example of a precipitation reaction? Explain.
3. **(4 pts)** Predict the products for the reaction that occurs when an aqueous solution of lead(II) nitrate ($\text{Pb}(\text{NO}_3)_2$) reacts chemically with an aqueous solution of potassium iodide (KI).
4. Write balanced chemical reactions for each of the following: **(14 pts)**
 - a) **(2 pts)** $\text{Al (s)} + \text{HNO}_3 \text{ (aq)} \rightarrow \text{Al}(\text{NO}_3)_3 \text{ (aq)} + \text{H}_2 \text{ (g)}$
 - b) **(4 pts)** Sodium sulfate + barium nitrate $\rightarrow$ barium sulfate + sodium nitrate
 - c) **(3 pts)** Calcium carbonate + heat $\rightarrow$ calcium oxide + carbon dioxide
 - d) **(3 pts)** Potassium + chlorine $\rightarrow$ potassium chloride
 - e) **(2 pts)** $\text{C}_{10}\text{H}_8 \text{ (s)} + \text{O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O} + \text{CO}_2$
5. **(5 pts)** During your laboratory experiment today you explored a precipitation reaction known as a double displacement reaction. A *double displacement* reaction has the following general formula: $\text{AB} + \text{CD} \rightarrow \text{AC} + \text{DB}$.

There are 4 other types of chemical reactions that are important for you to be able to identify in chemistry:

- 1) A single displacement reaction: $\text{AB} + \text{C} \rightarrow \text{AC} + \text{B}$
- 2) A synthesis reaction: $\text{A} + \text{B} \rightarrow \text{AB}$
- 3) A decomposition reaction: $\text{AB} + \text{heat} \rightarrow \text{A} + \text{B}$
- 4) A combustion reaction: $\text{C}_x\text{H}_y + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{CO}_2$

Use this information to help you classify each of the reactions in question 4 above.

Chemistry 101 Lab Manual Answers

Sasha Hu



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