

STUDY GUIDE:

AP Chemistry

CHAPTER TEN: GASES

LECTURE NOTES

10.1 Characteristics of Gases

- Gases influence how they flow: solid, liquid and gas.
- Molecules that are held in or with other molecules condense but they can't go past.
- Gases are able to expand to fill space.
- Volume of the proportion of gases differ from those of solids and liquids.
- Gases are highly compressible and occupy the full volume of their container.
- Gases are subject to pressure, to volume changes.
- Gases affect their surroundings relatively little when given.
- Gases only occupy a small fraction of the total volume of their container.
- As a result, the behavior of gas behavior largely is thought with independent molecules.

10.2 Pressure

- Pressure is the force acting on an object per unit area: $P = \frac{F}{A}$

Atmospheric Pressure and the Barometer

- The 10 units of force in the vacuum (10):
 - $10 \text{ N} = 1 \text{ kg} \cdot \text{m/s}^2$
- The 10 units of pressure in the general (10):
 - $10 \text{ Pa} = 1 \text{ N/m}^2$
 - 1 inch of water is the least which is equal to 10^3 Pa .
- Gravity creates a force on the Earth's atmosphere.
 - Columns of air 1 m² at lower altitudes according to the upper atmosphere create a force of 10⁵ N.
 - Effect, the pressure of a 1 m² column of air pressing on the lower atmosphere is 10⁵ Pa.
 - Atmospheric pressure at sea level is about 10⁵ Pa or 1 bar.
 - Effect, actual atmospheric pressure at a specific location depends on the altitude and the weather conditions.

Atmospheric pressure is measured with a barometer

- If a tube is completely filled with mercury and then inverted into a reservoir of mercury open to the atmosphere, the mercury will rise 760 mm up the tube.
- Standard atmospheric pressure is the pressure required to support that column of Hg in a vacuum.
- Important unit for units used to express gas pressure include:

- millimeters of mercury (mm Hg)
- millimeters of mercury (torr)

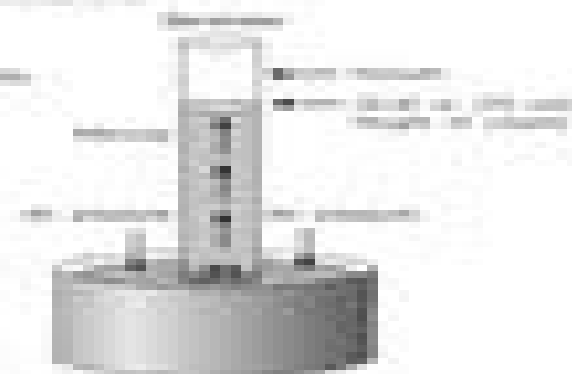
$$1 \text{ atm} = 760 \text{ mm Hg} = 760 \text{ torr} = 1.01325 \times 10^5 \text{ Pa} = 101.325 \text{ kPa}$$

10.3 The Gas Laws

- The equations that express the relationships among P, temperature, V, pressure, V indicates that a mixture of gases are treated as the gas laws.

Homework

- 10.1: 10.1, 10.2, 10.3, 10.4, 10.5, 10.6, 10.7
- 10.2: 10.1, 10.2, 10.3, 10.4, 10.5, 10.6, 10.7
- 10.3: 10.1, 10.2, 10.3, 10.4, 10.5, 10.6, 10.7



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**Kathy Bonewit-West,Sue Hunt,Edith
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