

Chemistry 111: Final Exam Topics

The final exam is comprehensive – covering the whole quarter. An approximate breakdown is as follows:

- Multiple Choice (160 points):
 - Roughly 2/3 based on older material (Topics from Exams 1-3)
 - Rest based on new material – Chapters 16, 17, 18
- Lab – related question (short answer= 40 points):
 - 1/5 (20 points) based on lab questions from Exams 1-3.
 - 1/5 (20 points) based on the Titration & Beer's Law labs.

Chapter 17: Net Ionic Equations

- "Electrolyte," how these solutions conduct electricity.
- Write total ionic equation, then cancel the "spectator ions" to get a net ionic equation
- Solubility Rules & precipitation reactions

Chapter 2

This is mostly science vocabulary!

- Scales in Science – Fig 2.1
- States (s/l/g)
- Review Goals – Page 36
- Review Vocabulary – Page 37

Chapter 3

- Scientific Notation
- Scientific Notation & the EE/EXP button on your calculator.
- Given factors (the table I gave you or given in the problem), find a unit path & do a conversion.
NOTE: you don't get the metric prefixes – memorize them.
- Metric Units for mass, volume, length
- Significant Figures
 - understand concept
 - Do the math (rules for + & -, × & ÷)
 - Know how to round
- Density = Mass ÷ Volume
- Convert °F ↔ °C ↔ K (equations might be given)

Chapter 18: Acid – Base Reactions

- Definition – Arrhenius & Bronsted (Not Lewis)
- Recognize Acids & Bases (from formula)
- Water as an acid/base (amphoteric)
- Water, K_w , and calculating $[H^+]$ & $[OH^-]$
- Calculate pH, pOH & relationship between them.
- Weak Acids & Bases, Conjugate Acids & Bases.

Chapter 10

- Electromagnetic spectrum (understand):
Continuous spectrum vs. line spectrum
- "Quantized"
- Quantum Model – slides from Monday are good...
- Valence electrons, ns^2np^6 notation
- Ionization Energy
- Vocabulary from end of chapter.

Chapter 11

- Monatomic Ions
- Ionic Bonds
- Covalent Bonds
- Covalent Bond Polarity
- Multiple Bonds
- Octet Rule Exceptions.

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