Name:

 Hydrogen, H₂, is used as rocket fuel. The hydrogen is burned in oxygen, O₂, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H2 is burned?

2 $H_2(g) + O_2(g) \rightarrow 2 H_2O(g)$; $\Delta H = -484 \text{ kJ}$

200.03 12 1 mol HZ = 99.2 mol Hz 99.2 mol Hz 1-484 KJ (2011)

Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide

 $CS_2(l) + 3O_2(g) \rightarrow CO_2(g) + 2SO_2(g) : \Delta H = -1077 \text{ kJ}$ What is ΔH for the following reaction?

charge 50 5:5 Not 1/2 CO2(8) + SO2(8) - 1/2 CS2(1) + 3/2 O2(8)

+ 1077 K3 = (538.5 K3)

3. Following is the equation for the combustion of acetone:

2 C₃H₆O(g) + 8 O₂(g) - 6 H₂O(g) + 6 CO₂(g) + 853.6 kcal/2

How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

- 853. 6 Kc-1 2 mol Catte a

4.5 mol (3400 = (-1920.6 KC-1)

 The specific heat of iron is 0.110 cal / g.°C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C? heat = 8 = m e AT = (5.2) (0.110 cml) (225-25'C) = (14.4)

 Hydrazine, N2H4, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the Grant process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.

> (1) $N_2(g) + 2H_2(g) \rightarrow$

 $N_2H_4(1) + O_2(g) \rightarrow N_2(g) + 2H_2O(1); \Delta H = -622.2 \text{ kJ}$

 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(f) : \Delta H = -285.8 \text{ kJ}$

N291+ 2420 (1) --> N2Hy (1) + O2(1, DH=622.2)

2H2G1 + OZG1 -> ZHZOQ1 DH= (Z)(-285.8) KJ NZG1 + ZHZG1 -> NZHYG1 (DH= + 50.6 KJ

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Mike Jess

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- 1. Understanding the eBook Chemistry 13 4 Practice Problems Answer Key
 - The Rise of Digital Reading Chemistry 13 4 Practice Problems Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry 13 4 Practice Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 13 4 Practice Problems Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry 13 4 Practice Problems Answer Key

- Personalized Recommendations
- Chemistry 13 4 Practice Problems Answer Key User Reviews and Ratings
- Chemistry 13 4 Practice Problems Answer Key and Bestseller Lists
- 5. Accessing Chemistry 13 4 Practice Problems Answer Key Free and Paid eBooks
 - Chemistry 13 4 Practice Problems Answer Key Public Domain eBooks
 - Chemistry 13 4 Practice Problems Answer Key eBook Subscription Services
 - Chemistry 13 4 Practice Problems Answer Key Budget-Friendly Options
- 6. Navigating Chemistry 13 4 Practice Problems Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chemistry 13 4 Practice Problems Answer Key Compatibility with Devices
 - Chemistry 13 4 Practice Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 13 4 Practice Problems Answer Key
 - Highlighting and Note-Taking Chemistry 13 4 Practice Problems Answer Key
 - Interactive Elements Chemistry 13 4 Practice Problems Answer Key
- 8. Staying Engaged with Chemistry 13 4 Practice Problems Answer Key
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 13 4 Practice Problems Answer Key
- 9. Balancing eBooks and Physical Books Chemistry 13 4 Practice Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 13 4 Practice Problems Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 13 4 Practice Problems Answer Key
 - Setting Reading Goals Chemistry 13 4 Practice Problems Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 13 4 Practice Problems Answer Key

- Fact-Checking eBook Content of Chemistry 13 4 Practice Problems Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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