

7. What is the oxidation state of the bold element in each compound below?

a) **Fe**Cl₂ _____

b) Na₂**S**O₃ _____

c) H₂**O**₂ _____

d) Na₃**P**O₄ _____

8. Place the following diatomic molecules in order of increasing polarity: CN, CF, O₂

9. Draw a plot of pressure vs. moles of gas for an ideal gas (label the axes, including units).

10. Circle the correct answer for each of the following:

a) the lowest electronegativity: F, Cl, **P**

b) would be most likely to form a stable anion with a -2 charge: Se, Cl, Mg

c) is diamagnetic: F, Fe³⁺, Na⁺

13. Like all carbonates, barium carbonate can be decomposed to the metal oxide and CO₂.



If the ΔH_{rxn} of barium carbonate decomposition is 269.3 kJ/mol, how many kJ are required to decompose 10.0 g of BaCO₃?

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Xiang Xie



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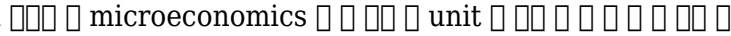
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