

Avogadro and the Mole Lab

NAME: _____ DATE: _____ PERIOD: _____

Background:

A mole is Avogadro's number of representative particles. Amedeo Avogadro determined that 6.02×10^{23} particles of any substance would equal 1 mole of that substance. The appropriate name for a representative particle is determined by the identity of the substance. A substance that is **ionic** (**metal-nonmetal**) would be appropriately named as **formula units**. A substance that is **molecular** (**nonmetal-nonmetal**) would be termed **molecules** and a substance that is an **element** would be labeled **atoms**. This number is too large to count so chemists use the concept of the mole to calculate the number of particles by obtaining the mass of the substance. This idea is referred to as "counting by weighing".

Grams → Moles → Representative Particles

Procedure:

- 1) From the front board, write down the average mass of an empty petri dish.
- 2) Go to each station and pick up the sample bottle with contents. Do not open the petri dish! Using a balance, obtain the mass of the petri dish and contents.
- 3) Using information written on the container label, fill in the data table below.

Data:

Sample Number	Chemical Formula	Chemical Name of Substance	Mass of empty petri dish	Mass of substance and petri dish
1				
2				
3				
4				
5				
6				
7				
8				
9				
10				

Give the number of the petri dish that contains a substance that would be used in each situation.

Plumbing _____ Flavoring Food _____ bath salts _____ For a headache _____
Computer chips _____ Edible Cooking Ingredients _____

Give the common name for the substances in the following numbered bottle.

Bottle #2 _____ Bottle # 4 _____ Bottle # 6 _____

Avogadro And The Mole Lab Answers

**United States. Air Force. Office of
Scientific Research**



Avogadro And The Mole Lab Answers:

Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations *Holt Chemistry* Ralph Thomas Myers, 2004 General Chemistry Brady, 1990-01-02 **Prep for Success in Chemistry, a Bridge Between Math and**

Science Laurie Sorge, 2011-01-28 Everything you need to succeed in Chemistry and may have missed along the way Cover

Fundamentals of General, Organic, and Biological Chemistry John R. Holum, 1982 User friendly this comprehensive and up to date edition is organized according to the principle that life in both health and disease has a molecular basis Carefully worked example exercises enhanced by Analysis and Check sections enable students to examine and think about a problem not just answer it Includes numerous environmental applications Campbell Biology Australian and New Zealand Edition

Jane B. Reece, Noel Meyers, Lisa A. Urry, Michael L. Cain, Steven A. Wasserman, Peter V. Minorsky, 2015-05-20 Over nine successful editions CAMPBELL BIOLOGY has been recognised as the world's leading introductory biology textbook The Australian edition of CAMPBELL BIOLOGY continues to engage students with its dynamic coverage of the essential elements of this critical discipline It is the only biology text and media product that helps students to make connections across different core topics in biology between text and visuals between global and Australian New Zealand biology and from scientific study to the real world The Tenth Edition of Australian CAMPBELL BIOLOGY helps launch students to success in biology through its clear and engaging narrative superior pedagogy and innovative use of art and photos to promote student learning It continues to engage students with its dynamic coverage of the essential elements of this critical discipline This Tenth Edition with an increased focus on evolution ensures students receive the most up to date accurate and relevant information *Educational Tests and Measurement* Anthony J. Nitko, 1983 *Contemporary Chemistry* Leonard

Saland, 1993 This comprehensive guide gives you lesson plans activities and tests for two sequential semester long chemistry courses It is designed to work with our student book Contemporary Chemistry Each lesson plan features a DO NOW section to engage students as soon as they get to class instructional objectives an aim for that class period a motivational application questions or demonstrations to help students draw valid conclusions homework assignments You also get term calendars weekly tests and complete answer keys **Micro Experiments** Holt Rinehart & Winston, Holt, Rinehart and Winston

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Chemistry, Problem-Solving Worktext James E. Brady,Frederick Senese,2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet *Elements*

of General and Biological Chemistry John R. Holum, 1983 *The Measurement of Inclusive Lambda, K-short and Lambda Bar Production and Lambda Polarization from Interactions of 400 GeV/c Protons with Hydrogen* Ronald Allan Grobel, 1980

Chemistry James E. Brady, John R. Holum, 1995-12-29 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping readers develop skills in analytical thinking and problems solving **The**
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