Gen Chem 2 LG Session 2

Answer Key

Question 1: A 1.12 M solution of citric acid (H3C6H5O7) in water has a density of 1.08 g/cm3. Calculate the mass percent, molality, and mole fraction of the citric acid.

Assume 1 L of Solution : 1.12 moles (the Acid & mass of Molar Mass is 192,1233 Solution is 1080 grams

Mass is 192,1233 Mass 8 = 315.178 grams

Mass of 1420 = 1089 - 215.178 grams

Molality = 1.12 moles = 1.3 mol mg

Molality = 1.12 moles = 1.3 moles = 1.12 moles = 1.12 mg

Question 2: Pure solvent X boils at 81.1°C. When 11.00 g of solute A (non-electrolyte, MM = 88.1 g/mol) is dissolved in 150.0 g of solvent X, the new boiling point is 83.5°C. When 4.00 g of solute B (non-electrolyte) is dissolved in 50.0 g of solvent X, the solution boils at 82.4°C. What is the molar mass of solute B?

A:
$$\Delta B \hat{\rho}^{2} = 2.4 \,^{\circ}C$$
 $VanHfff = 1$
 $Molality = \frac{Moles}{Mg \, Solvent} = \frac{(19/88.19/Mol)}{0.150 \, Mg} = 0.8334 \, \frac{Mol}{Mg}$
 $2.4 = \Delta B \hat{\rho} = 1.0.8534 \, \frac{M}{Mg} \cdot \frac{M}{Mg}$
 $Value = \frac{2.4 \,^{\circ}C}{0.8534 \, \frac{M}{Mg}} \cdot \frac{M}{Mg}$
 $Value = \frac{2.4 \,^{\circ}C}{0.8534 \, \frac{M}{Mg}} \cdot \frac{M}{Mg}$
 $Value = \frac{2.4 \,^{\circ}C}{0.8534 \, \frac{M}{Mg}} \cdot \frac{M}{Mg}$
 $Value = \frac{2.4 \,^{\circ}C}{0.85334 \, \frac{M}{Mg}} \cdot \frac{M}{Mg}$
 $Value = \frac{1.3 \,^{\circ}C}{0.85833} = 0.45088 \, \frac{M}{Mg}$
 $Value = \frac{M}{Mg} = \frac{M}{Mg} \cdot \frac{M}{Mg}$

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