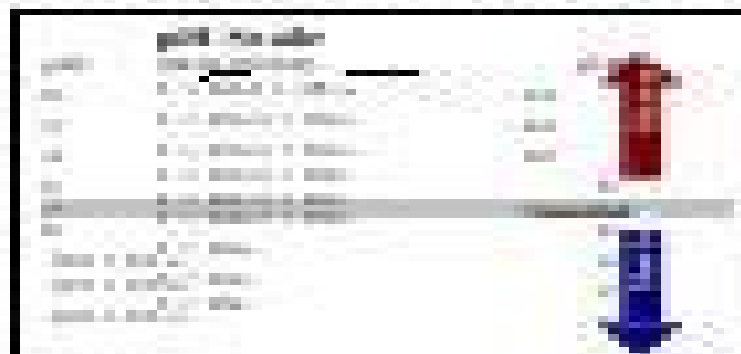


# The pH Scale

Chem. 100 Lecture 10p 2

Name: \_\_\_\_\_

An acidic solution has a pH less than 7. A basic solution has a pH greater than 7. A neutral solution has a pH of 7. The pH scale is logarithmic, meaning that the value decreases by a factor of 10 for each unit increase in pH. For example, a solution with a pH of 3 is 10 times more acidic than a solution with a pH of 4, and 100 times more acidic than a solution with a pH of 5.



The pH scale is logarithmic, meaning that the value decreases by a factor of 10 for each unit increase in pH. For example, a solution with a pH of 3 is 10 times more acidic than a solution with a pH of 4, and 100 times more acidic than a solution with a pH of 5.

## Examples

What is the pH of a solution with  $[H^+] = 1 \times 10^{-3} M$ ?

$pH = -\log [H^+]$   
 $pH = -\log (1 \times 10^{-3})$   
 $pH = 3$

What is the pH of a solution with  $[OH^-] = 1 \times 10^{-4} M$ ?

$pH = 14 - pOH$   
 $pOH = -\log [OH^-]$

Since  $[OH^-] = 1 \times 10^{-4} M$ ,  $pOH = 4$ .  
 $pH = 14 - 4 = 10$

What is the pH of a solution with  $[H^+] = 1 \times 10^{-10} M$ ?

$pH = -\log [H^+]$   
 $pH = -\log (1 \times 10^{-10})$   
 $pH = 10$

What is the pH of a solution with  $[OH^-] = 1 \times 10^{-5} M$ ?

$pH = 14 - pOH$   
 $pOH = -\log [OH^-]$   
 $pOH = -\log (1 \times 10^{-5}) = 5$   
 $pH = 14 - 5 = 9$

Rank the following solutions from most acidic to most basic:

1. A solution with a pH of 3
2. A solution with a pH of 10
3. A solution with a pH of 7
4. A solution with a pH of 1
5. A solution with a pH of 12
6. A solution with a pH of 5
7. A solution with a pH of 9
8. A solution with a pH of 6
9. A solution with a pH of 8
10. A solution with a pH of 4
11. A solution with a pH of 11
12. A solution with a pH of 2
13. A solution with a pH of 13
14. A solution with a pH of 0
15. A solution with a pH of 14

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# Chem Worksheet 19 3 Answers

**Eugene Joseph Connelly**



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