

Class Activity 24

The Ionic Bond

Model 1: A Cation and an Anion Form an Ionic Bond

Recall that, as ions form in chemical reactions, they usually release (or absorb) substantial amounts of energy as well as many electrons are produced, ions (charged species) are produced and are attracted to a species of the opposite charge. In some cases, the change in ion ionization energy is sufficient to understand the "electron-gaining" of the cation and the electron "loss" of the anion. In some cases, the energy of the cation and the anion is not sufficient to understand the "electron-gaining" of the cation and the electron "loss" of the anion. In some cases, the energy of the cation and the anion is not sufficient to understand the "electron-gaining" of the cation and the electron "loss" of the anion.

Critical Thinking Questions

1. The ionization energy of sodium is 496 kJ/mol. The electron affinity of chlorine is 349 kJ/mol. Is the reaction $\text{Na} + \text{Cl} \rightarrow \text{Na}^+ + \text{Cl}^-$ exothermic or endothermic? Explain this result.

Exothermic. The ionization energy of sodium is 496 kJ/mol and the electron affinity of chlorine is 349 kJ/mol. The reaction is exothermic because the energy released by the ionization of sodium is greater than the energy absorbed by the electron affinity of chlorine.

2. The ionization energy of sodium is 496 kJ/mol. The electron affinity of chlorine is 349 kJ/mol. Is the reaction $\text{Na} + \text{Cl} \rightarrow \text{Na}^+ + \text{Cl}^-$ exothermic or endothermic? Explain this result.

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Exothermic. The ionization energy of sodium is 496 kJ/mol and the electron affinity of chlorine is 349 kJ/mol. The reaction is exothermic because the energy released by the ionization of sodium is greater than the energy absorbed by the electron affinity of chlorine.

Exercises

1. Write the electron configuration for Na^+ and Cl^- and explain the result.
2. Write the electron configuration for Na^+ and Cl^- and explain the result.

Na: $1s^2 2s^2 2p^6 3s^1$ Na⁺: $1s^2 2s^2 2p^6$ Cl: $1s^2 2s^2 2p^6 3s^2 3p^5$ Cl⁻: $1s^2 2s^2 2p^6 3s^2 3p^6$

Na: $1s^2 2s^2 2p^6 3s^1$ Na⁺: $1s^2 2s^2 2p^6$ Cl: $1s^2 2s^2 2p^6 3s^2 3p^5$ Cl⁻: $1s^2 2s^2 2p^6 3s^2 3p^6$

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Richard Samuel Moog



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