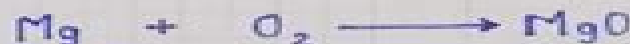


Chemical Reactions

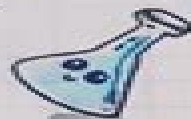
Chemical Equation

Reactants \longrightarrow Products
Left hand side Right hand side

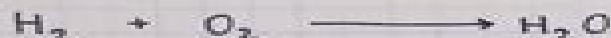
magnesium + Oxygen \longrightarrow magnesium oxide



BALANCING EQUATIONS



number of atoms of every element before and after reaction is same



Chemical equations are a way of showing chemical reactions.

LAW OF CONSERVATION OF MATTER
a chemical equation must be balanced.

Chemical reactions

COMBINATION

Two or more substance combine to form a single product.



DECOMPOSITION

Single reactant breaks up to form two or more products



DISPLACEMENT

More reactive element removes the other element from the compound and itself forms the compound



- Chemical reactions can occur only:
 \hookrightarrow when the chemical change.

Rusting of iron, Burning of coal.

DOUBLE DISPLACEMENT

Exchange of ions between the reactants take place.



REDUCTION

gain of hydrogen or loss of oxygen.

redox

OXIDATION

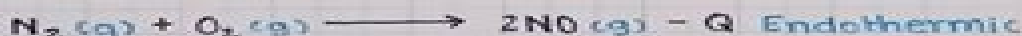
gain of oxygen or loss of hydrogen.



PHYSICAL STATE -



CHANGE IN HEAT -



EVOLUTION OF GAS -



PRECIPITATION -



DIFF. REACTIONS CONDITION -



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Michael P. Olmsted



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