

# CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS

## FOLLOW-UP PROBLEMS

17.1A **Plan:** First, balance the equations and then write the reaction quotient. Products appear in the numerator of the reaction quotient and reactants appear in the denominator; coefficients in the balanced reaction become exponents.

**Solution:**



$$\text{Reaction quotient: } Q_c = \frac{[\text{NO}]^4 [\text{H}_2\text{O}]^6}{[\text{NH}_3]^4 [\text{O}_2]^5}$$



$$\text{Reaction quotient: } Q_c = \frac{[\text{NH}_3]^2}{[\text{NO}_2]^2 [\text{H}_2]^7}$$



$$\text{Reaction quotient: } Q_c = [\text{O}_2]^3$$

17.1B **Plan:** First, balance the equations and then write the reaction quotient. Products appear in the numerator of the reaction quotient and reactants appear in the denominator; coefficients in the balanced reaction become exponents.

**Solution:**



$$\text{Reaction quotient: } Q_c = \frac{[\text{CO}]^2 [\text{H}_2]^2}{[\text{CH}_4] [\text{CO}_2]}$$



$$\text{Reaction quotient: } Q_c = \frac{[\text{H}_2\text{O}]^2}{[\text{H}_2\text{S}]^2 [\text{SO}_2]}$$



$$\text{Reaction quotient: } Q_c = \frac{[\text{NaCN}]}{[\text{HCN}] [\text{NaOH}]}$$

17.2A **Plan:** Add the individual steps to find the overall equation, canceling substances that appear on both sides of the equation. Write the reaction quotient for each step and the overall equation. Multiply the reaction quotients for each step and cancel terms to obtain the overall reaction quotient.

**Solution:**



Canceling the reactants leaves the overall equation as  $\text{Br}_2(g) + \text{H}_2(g) \rightleftharpoons 2\text{HBr}(g)$ .

Write the reaction quotients for each step:

$$Q_{c1} = \frac{[\text{Br}]^2}{[\text{Br}_2]} \quad Q_{c2} = \frac{[\text{HBr}][\text{H}]}{[\text{Br}][\text{H}_2]} \quad Q_{c3} = \frac{[\text{HBr}]}{[\text{H}][\text{Br}]}$$

and for the overall equation:

$$Q_c = \frac{[\text{HBr}]^2}{[\text{H}_2][\text{Br}_2]}$$

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1. Understanding the eBook Chemical Equilibrium Solutions Manual Answers
  - The Rise of Digital Reading Chemical Equilibrium Solutions Manual Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Solutions Manual Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Equilibrium Solutions Manual Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Solutions Manual Answers
  - Personalized Recommendations
  - Chemical Equilibrium Solutions Manual Answers User Reviews and Ratings

- Chemical Equilibrium Solutions Manual Answers and Bestseller Lists
- 5. Accessing Chemical Equilibrium Solutions Manual Answers Free and Paid eBooks
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  - Chemical Equilibrium Solutions Manual Answers Budget-Friendly Options
- 6. Navigating Chemical Equilibrium Solutions Manual Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Equilibrium Solutions Manual Answers Compatibility with Devices
  - Chemical Equilibrium Solutions Manual Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Equilibrium Solutions Manual Answers
  - Highlighting and Note-Taking Chemical Equilibrium Solutions Manual Answers
  - Interactive Elements Chemical Equilibrium Solutions Manual Answers
- 8. Staying Engaged with Chemical Equilibrium Solutions Manual Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Equilibrium Solutions Manual Answers
- 9. Balancing eBooks and Physical Books Chemical Equilibrium Solutions Manual Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Equilibrium Solutions Manual Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Equilibrium Solutions Manual Answers
  - Setting Reading Goals Chemical Equilibrium Solutions Manual Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Equilibrium Solutions Manual Answers
  - Fact-Checking eBook Content of Chemical Equilibrium Solutions Manual Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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