

Chemistry--Unit 10: Chemical Periodicity
Practice Problems

I. Classification of the elements

- 1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$ or abbreviated form*: $[\text{Ne}] 3s^2 3p^2$ *may use this form now!!!

- 2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$ or $[\text{Kr}] 5s^2 4d^{10} 5p^5$

- 3) Which family of elements is characterized by an $s^2 p^3$ configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a $4s^2 4p^5$ electron configuration **Br, bromine**
- the 4th period Group 16 element **Se, selenium**

- 5) Identify the elements that have the following outer electron configurations?

- $2s^2 2p^4$ **O, oxygen**
- $4s^2$ **Ca, Calcium**
- $4s^2 3d^{10}$ **Zn, zinc**

- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

- 7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

- 8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

- 10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

Chemical Periodicity Practice Problems Answers

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The Problem of Physico-chemical Periodicity Ernest Sydney Hedges, James Eckersley Myers, 1926 **GO TO Objective**
NEET 2021 Chemistry Guide 8th Edition Disha Experts, *Student Solutions Manual* H. Stephen Stoker, 1999-02-11

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presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

Introduction to Chemical Principles Howard Stephen Stoker,1990 Introduction to Chemical Principles is a text for students who have had little to no previous instruction in chemistry or who had such instruction long enough ago that a thorough review is needed preface

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg,2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change

McGraw-Hill's SAT Subject Test: Chemistry, 2ed Thomas Evangelist,2009-02-01 We want to help you score high on the SAT Chemistry test We ve put all of our proven expertise into McGraw Hill s SAT Subject Test Chemistry to make sure you re fully prepared for this difficult exam With this book you ll get essential skill building techniques and strategies created by a leading high school chemistry teacher You ll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill s SAT Subject Test Chemistry we ll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score

Problem Exercises for General Chemistry G.Gilbert Long,Forrest C. Hentz,1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT

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Problem Exercises for General Chemistry George Gilbert Long,Forrest C. Hentz,1982

Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill,2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change

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