

CHEMICAL QUANTITIES

Chapter Quiz

Fill in the word(s) that will make each statement true.

1. Counting, finding mass, and finding volume are all ways of measuring the amount of 1.
2. The 2 is the mass, in amu, of one atom of a monatomic element.
3. The 3 is the mass in grams of one mole of a substance.
4. The representative particle of most elements is the 4.
5. To convert moles of oxygen gas (O_2) to mass in grams, you can use the conversion factor 5.

1.

101

2.

101

3.

102

4.

102

5.

102

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 6. The density of ethane gas, C_2H_6 , at STP is 1.34 g/L. 102
- _____ 7. The volume of 3 moles of oxygen at STP is 67.2 L. 102
- _____ 8. One mole of a gas occupies 22.4 L. 102
- _____ 9. To convert a volume of 3.20 L of hydrogen gas to the number of moles, 102
you would use the conversion factor $\frac{22.4\text{ L}}{1.00\text{ mol H}_2}$.
- _____ 10. The density of a gas at STP can be used to calculate the molar mass of that gas. 102

Solve the following problems. Show your work in the space provided.

11. What mass of gold contains as many atoms as 3.05 g of silver? 102
12. What is the empirical formula of a compound that is 40.7% carbon, 54.2% oxygen, and 5.1% hydrogen? 103

Chemical Quantities Chapter Quiz Answer Key

**Gerald Karp, Janet Iwasa, Wallace
Marshall**



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General, Organic, & Biological Chemistry Janice Gorzynski Smith,2024-12 My goal in writing this text was to relate the fundamental concepts of general organic and biological chemistry to the world around us and in this way illustrate how chemistry explains many aspects of everyday life A key feature is the use of molecular art to illustrate and explain common phenomena we encounter every day Each topic is broken down into small chunks of information that are more manageable and easily learned Students are given enough detail to understand basic concepts such as how soap cleans away dirt and why trans fats are undesirable in the diet without being overwhelmed

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