

Chemistry 12
Worksheet 2-2
LeChatelier's Principle

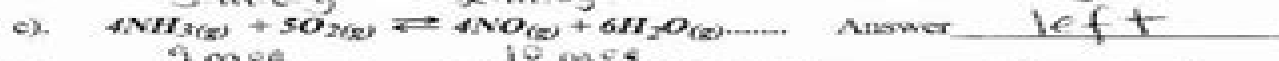
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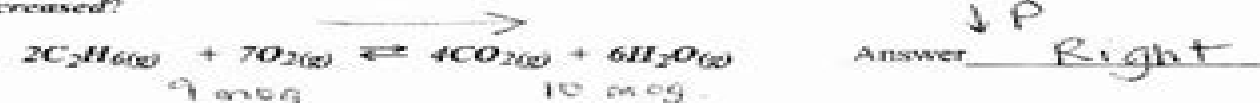
1. In order to decide what effect a *change in total pressure* will have on an equilibrium system with gases, what is the first thing you should do when given the balanced equation?

Add up the moles of gas (m.o.g) on both sides of the equation

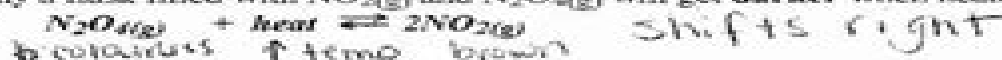
2. Predict which way the following equilibrium systems will shift when the *total pressure* is *increased*. (NOTE: Some may have no shift)



3. Which way will the following equilibrium shift if the *total pressure* on the system is *decreased*?



4. Explain why a flask filled with $\text{NO}_2(\text{g})$ and $\text{N}_2\text{O}_4(\text{g})$ will get darker when heated. Use the equation: $\text{N}_2\text{O}_4(\text{g}) + \text{heat} \rightleftharpoons 2\text{NO}_2(\text{g})$ *shifts right*



Heating causes the equilibrium to shift right which will produce more NO_2 gas. This is brown so flask will show a darker colour.

5. State *Le Chatelier's Principle*.

When a stress is applied to a system at equilibrium, the equilibrium will SHIFT to counteract the imposed stress

Chemistry 12 Worksheet 2 2 Answers

Kenneth T. Henson



Chemistry 12 Worksheet 2 2 Answers:

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