

UNIT	Reading	Vocabulary	Grammar
1 It's personal! pages 5-16	missing sentences • identifying language links in the text and the sentences	• words related to people • collocations & expressions, phrasal verbs, word formation	• present simple, present continuous, stative verbs • relative clauses
2 Life's Great Mysteries pages 17-28	multiple choice • answering a question in your own words to avoid being distracted by the answer choices	• words related to mysteries • phrasal verbs, commonly confused words, collocations & expressions	• past simple, past continuous • used to, would, be/get used to
Progress Review 1 pages 29-30			
3 The World of Sport pages 31-42	multiple matching • summarising the paragraphs for a better understanding of the text	• words related to sport • word formation, collocations & expressions, phrasal verbs	• present perfect simple, present perfect continuous • past simple, present perfect simple
4 Shopping Therapy pages 43-54	multiple choice • answering questions about four related texts	• words related to shopping • phrasal verbs, prepositions, commonly confused words	• past perfect simple, past perfect continuous • past perfect tenses, past simple
Progress Review 2 pages 55-56			
5 The Environment pages 57-68	multiple matching • identifying key words in the questions and similar ideas in the texts	• words related to the environment • word formation, prepositions, phrasal verbs	• future simple, <i>shall</i> , <i>be going to</i> , future continuous • future perfect simple, future perfect continuous, temporals
6 Crime pages 69-80	missing sentences • making sure sentences fit logically with what is before and after each gap	• words related to crime, and punishment • phrasal verbs, commonly confused words, collocations & expressions	• modals: ability, permission, requests, offers & suggestions • modals: obligation, necessity, prohibition, possibility, deduction, advice, criticism
Progress Review 3 pages 81-82			
7 The Media pages 83-94	multiple choice • eliminating obviously wrong answer options	• words related to different kinds of media • prepositions, collocations & expressions, word formation	• the passive: tenses, gerunds, infinitives & modals • the passive: impersonal and personal structures
8 A Hard Day's Work pages 95-106	multiple choice • identifying different text types	• words related to jobs and work • commonly confused words, word formation, collocations & expressions	• conditionals: zero, first, second & third • mixed conditionals
Progress Review 4 pages 107-108			
9 Getting Away pages 109-120	multiple matching • scanning the texts for specific information	• words related to holiday and travel • commonly confused words, word formation, prepositions	• gerunds & infinitives, <i>make & let</i> • comparison of adjectives & adverbs
10 Where on earth? pages 121-132	multiple choice • answering questions about what dates refer to	• words related to the world, language and culture • commonly confused words, word formation, prepositions	• reported speech: statements, questions, commands & requests • reported speech: reporting verbs
Progress Review 5 pages 133-134			
11 High-Tech World pages 135-146	missing sentences • identifying sentences that explain something in the text	• words related to technology • prepositions, phrasal verbs, commonly confused words	• causative forms • wishes & preferences
12 Happy to help! pages 147-158	multiple choice • understanding the text as a whole	• words related to services and helping people • phrasal verbs, word formation, collocations & expressions	• question tags, pronouns • <i>not only ... but also</i> , <i>had better</i> , <i>it's (about/high) time</i>
Progress Review 6 pages 159-160			

Grammar Reference	pages 161-177
Irregular Verbs	pages 178-179
Writing Bank	pages 180-186
Speaking Bank	page 187
Speaking Information	pages 188-191

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Instructor's Guide and Answer Keys to Accompany The Holt Handbook ,2002

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^-aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation $4\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$ because two NH_4^+aq

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