

## Unit 5 Problem Set Three: Limiting Reagents and Percent Yield

For each problem, show your calculations which should include properly units, solutions. The final answer should be indicated with proper significant figures, unit and calculation. Answers only are not accepted.

1. In reaction 1, a theoretical yield of 118.3 g  $\text{NF}_3$ , but only 11.87 g  $\text{NF}_3$  are obtained in the lab, what is the percent yield of  $\text{NF}_3$  for this reaction?

$$11.87 \text{ g NF}_3 / 118.3 \text{ g NF}_3 \times 100 = 10.03\%$$

2. Aluminum can be synthesized by the following reaction:  $2\text{Al}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Al} + 3\text{CO}_2$

If the following reaction occurred with 5.00g of Hydrogen gas and 11.7g of Fluorine gas, what is the theoretical yield of ammonia?

$$5.00 \text{ g H}_2 \times 1 \text{ mol H}_2 / 2.02 \text{ g} \times 2 \text{ mol NH}_3 / 2 \text{ mol H}_2 \times 17.03 \text{ g NH}_3 / 1 \text{ mol NH}_3 = 53.5 \text{ g NH}_3$$

$$11.7 \text{ g F}_2 \times 1 \text{ mol F}_2 / 38.00 \text{ g} \times 2 \text{ mol NH}_3 / 3 \text{ mol F}_2 \times 17.03 \text{ g NH}_3 / 1 \text{ mol NH}_3 = 10.7 \text{ g NH}_3$$

$$10.7 \text{ g NH}_3 < 53.5 \text{ g NH}_3 \rightarrow 10.7 \text{ g NH}_3$$

$$10.7 \text{ g NH}_3$$

$$10.7 \text{ g NH}_3 / 53.5 \text{ g NH}_3 \times 100 = 20.0\%$$

3. Given the following reactions:  $\text{Sn}_3(\text{PO}_4)_4 + 3\text{H}_2\text{CO}_3 \rightarrow \text{Sn}_3(\text{CO}_3)_4 + 4\text{H}_3\text{PO}_4$

- a. Balance the reaction:  $\text{Sn}_3(\text{PO}_4)_4 + 3\text{H}_2\text{CO}_3 \rightarrow \text{Sn}_3(\text{CO}_3)_4 + 4\text{H}_3\text{PO}_4$

- b. If 10.0 g/mol of the triphosphate react with an excess of carbonic carbonate, what is the theoretical yield of the  $\text{P}_2\text{O}_5$  carbonate?

$$10.0 \text{ g Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol Sn}_3(\text{PO}_4)_4 / 758.6 \text{ g} \times 3 \text{ mol H}_2\text{CO}_3 / 1 \text{ mol Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol P}_2\text{O}_5 / 1 \text{ mol H}_2\text{CO}_3 \times 108.0 \text{ g P}_2\text{O}_5 / 1 \text{ mol P}_2\text{O}_5 = 3.95 \text{ g P}_2\text{O}_5$$

$$10.0 \text{ g Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol Sn}_3(\text{PO}_4)_4 / 758.6 \text{ g} \times 3 \text{ mol H}_2\text{CO}_3 / 1 \text{ mol Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol P}_2\text{O}_5 / 1 \text{ mol H}_2\text{CO}_3 \times 108.0 \text{ g P}_2\text{O}_5 / 1 \text{ mol P}_2\text{O}_5 = 3.95 \text{ g P}_2\text{O}_5$$

- a. If the reaction produced 20.75 g/mol of the  $\text{P}_2\text{O}_5$  carbonate, what was the actual yield? 20.75 g/mol of the  $\text{P}_2\text{O}_5$  carbonate

- b. If the reaction used 10.0 g/mol of  $\text{Sn}_3(\text{PO}_4)_4$ , what is the theoretical yield of carbonic phosphate?

$$10.0 \text{ g Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol Sn}_3(\text{PO}_4)_4 / 758.6 \text{ g} \times 3 \text{ mol H}_2\text{CO}_3 / 1 \text{ mol Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol P}_2\text{O}_5 / 1 \text{ mol H}_2\text{CO}_3 \times 108.0 \text{ g P}_2\text{O}_5 / 1 \text{ mol P}_2\text{O}_5 = 3.95 \text{ g P}_2\text{O}_5$$

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- a. If the reaction reacted to 10.0 g/mol, what was the actual yield?

$$10.0 \text{ g Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol Sn}_3(\text{PO}_4)_4 / 758.6 \text{ g} \times 3 \text{ mol H}_2\text{CO}_3 / 1 \text{ mol Sn}_3(\text{PO}_4)_4 \times 1 \text{ mol P}_2\text{O}_5 / 1 \text{ mol H}_2\text{CO}_3 \times 108.0 \text{ g P}_2\text{O}_5 / 1 \text{ mol P}_2\text{O}_5 = 3.95 \text{ g P}_2\text{O}_5$$

4. Given the following reactions:  $\text{PbCl}_2 + \text{Na}_2\text{PO}_4 \rightarrow \text{Pb}_3(\text{PO}_4)_2 + \text{NaCl}$

- a. Balance the reaction:  $3\text{PbCl}_2 + 2\text{Na}_2\text{PO}_4 \rightarrow \text{Pb}_3(\text{PO}_4)_2 + 6\text{NaCl}$

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**David Baud**



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