

Released AP Multiple Choice Questions – Acids & Bases / Aqueous Equilibrium

D	33. The pH of 0.1-molar ammonia is approximately (A) 1 (B) 4 (C) 7 (D) 11 (E) 14
E	48. Which of the following ions is the strongest Lewis acid? (A) Na^+ (B) Cl^- (C) CH_3COO^- (D) Mg^{2+} (E) Al^{3+}
C	49. Each of the following can act as both a Brønsted acid and a Brønsted base EXCEPT (A) HCO_3^- (B) H_2PO_4^- (C) NH_4^+ (D) H_2O (E) HS^-
B	53. Which, if any, of the following species is in the greatest concentration in a 0.100-molar solution of H_2SO_4 in water? (A) H_2SO_4 molecules (B) H_3O^+ ions (C) HSO_4^- ions (D) SO_4^{2-} ions (E) All species are in equilibrium and therefore have the same concentrations.
E	71. Which of the following reactions does NOT proceed significantly to the right in aqueous solutions? (A) $\text{H}_3\text{O}^+ + \text{OH}^- \rightarrow 2 \text{H}_2\text{O}$ (B) $\text{HCN} + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{CN}^-$ (C) $\text{Cu}(\text{H}_2\text{O})_4^{2+} + 4 \text{NH}_3 \rightarrow \text{Cu}(\text{NH}_3)_4^{2+} + 4 \text{H}_2\text{O}$ (D) $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$ (E) $\text{H}_2\text{O} + \text{HSO}_4^- \rightarrow \text{H}_2\text{SO}_4 + \text{OH}^-$
E	75. If the acid dissociation constant, K_a , for an acid HA is 8×10^{-4} at 25 °C, what percent of the acid is dissociated in a 0.50-molar solution of HA at 25 °C? (A) 0.08% (B) 0.2% (C) 1% (D) 2% (E) 4%
A	34. All of the following species can function as Brønsted-Lowry bases in solution EXCEPT (A) H_2O (B) NH_3 (C) S^{2-} (D) NH_4^+ (E) HCO_3^-

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Mina Lebitz



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