

COLLIGATIVE PROPERTIES OF SOLUTIONS

A colligative property of a solution is a property that is only affected by the number of solute particles. There are four colligative properties of solutions, viz :



Vapor Pressure Drop (ΔP)

When there is a lid on the container, the gas phase molecules are trapped, they are a vapor. The vapor creates a pressure



If a solute is added to a solvent, the solute particles will prevent the solvent particles from turning into vapor. This causes a decrease in vapor pressure (ΔP).

Lots of dissolved substances $\rightarrow \Delta P$ big \rightarrow low vapor pressure (P).

The amount of vapor pressure drop can be determined by the following formula

a. Non-electrolyte solution

- $$\Delta P = P^* - P$$
$$\Delta P = X_t \times P^*$$
$$\Delta P = \frac{n_t}{n_t + n_p} \times P^*$$

- $$P = X_p \times P^*$$
$$P = \frac{n_p}{n_t + n_p} \times P^*$$

b. Electrolyte Solution

The electrolyte solution in the solvent can ionize. This results in the electrolyte solution having a more significant number of particles than the non-electrolyte solution at the same concentration. The degree of ionization is a direct measure of its ionizing state. The magnitude of the degree of ionization is expressed as :

Colligative Properties Of Solutions Study Guide Answers

Liying Dong



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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 H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Denniston Katherine, 2006 [Study Guide for Chemical Principles](#) [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text. The new edition of the Study Guide includes many new worked out examples.

Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04 [Answers to Study Questions Oceanography](#) Garrison, 2001-07

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