

CHEMISTRY 1 – SEMESTER 2

REVIEW GUIDE

Chemical Reactions

Chemical Reaction: process by which one or more substances are changed into different substances.

Total mass of reactants must equal total mass of products

Evidence of reactions:

- Heat
- Light
- Gas
- Precipitate
- Color change

Reactants: substances being reacted

Products: substances produced by reaction

Chemical Equation: a representation of a chemical reaction using symbols and formulas

$\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$ Reaction between methane and oxygen to produce carbon dioxide and water

Coefficients: relative number of moles

1 mole of CH_4 + 2 moles of O_2 1 mole CO_2 + 2 moles H_2O

$3\text{Fe}(\text{s}) + 4\text{H}_2\text{O}(\text{g}) \rightarrow \text{Fe}_3\text{O}_4(\text{s}) + 4\text{H}_2(\text{g})$

States of matter: solid/liquid/gas

Solid iron + gas water solid iron oxide + gas hydrogen

Word Equations: chemical reactions represented in words without formulas or symbols (or coefficients)

Sodium oxide + water → sodium hydroxide

$\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{NaOH}$ Unbalanced

Balanced Equations:

- Correct formulas
- Correct number of moles

Never change subscripts

Polyatomic ions are single units

Count all atoms at the end to make sure

Unbalanced: $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

Unbalanced: $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

Balanced: $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$

Balanced: $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$

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**Katherine J. Denniston, Denniston
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