



Covalent Compounds Unit Study Guide

Name _____
 Teacher _____
 Date: _____
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Reference:

Chemistry in the Community, pp 181-183,
 Notes
 Naming covalent compounds worksheet
 Viscosity Lab
 Exploring Properties of Ionic and Covalent compounds

Vocabulary

Understand and apply the definitions of the following terms:

| | |
|-------------------------------------|----------------|
| Covalent | Octet rule |
| Valence electrons | Polymers |
| e dot diagrams (Lewis dot diagrams) | Viscosity |
| Polar | Boiling points |
| Nonpolar | |
| Intermolecular forces | |

Targeted Skills

- Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
- Know the following properties:
 - > Ionic compounds = the transfer of electron(s), i.e. "gain or lose"
 Metal cations & nonmetal anion (or polyatomic ion)
 Net charge = zero (criss cross charges) $X^{+n}Y^{-m} = X_nY_m$
 Dissociate (dissolve) in water
 - > Molecular compounds = sharing electrons by forming a covalent bond
 2 or more nonmetals
 Prefixes indicate # of each atom
 Polar molecules have a δ^+ & a δ^-
 Polar molecules dissolve in water
 Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.
 - > Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
- Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
- Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
- Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
- Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

Make sure you remember from past tests?

| | |
|----------------------|--------------------------------|
| Atomic structure | Solubility |
| Significant figures | Safety |
| Accuracy & precision | Elements |
| Mixtures | Ionic compound formula writing |
| Nomenclature | |

Covalent Compounds Study Guide

Susan McMurry



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