

14: Chemical Bonding Theories

Key Chemistry Terms

- Ionic Bond:** Bond formed from electrostatic attraction between ions (charged atoms)—formed from metals with nonmetals.
- Covalent Bond:** A bond formed between two nonmetals that involves shared electrons.
- Polar covalent bond:** Bond formed between two nonmetals—uneven sharing of electrons.
- Metallic bond:** Formed between metal atoms—electrons are pooled between the network of atoms.
- Polar bond:** Bond between two atoms with a great difference in electronegativities—uneven sharing of electrons resulting in a slightly positive region and a slightly negative region of the bond.
- Electronegativity:** An atom's "pull" on electrons shared with another atom.
- Polar molecules:** Molecule where the polar bonds do not cancel each other out in 3D orientation.
- Isomers:** Molecules with the same chemical formula, but different bonding structure.
- Resonance:** Two compounds with multiple bonds in different locations.
- Valence Bond Theory:** Overlap of atomic orbitals form bond.
- Sigma (σ) bonds:** First bond between two atoms, formed from head on overlap of orbitals.
- Pi (π) bonds:** 2nd or 3rd bond between two atoms formed from overlap of parallel p orbitals.
- Hybridization:** Atomic orbitals hybridize into orbitals of the same energy for all sigma bonds of the atom.

Characteristics of bond types

Bond type	Happens between	Electrons are
Ionic	Metal & non-metal	Transferred
Covalent	Non-metals	Shared
Polar Covalent	Non-metals	Shared unevenly
Metallic	Metals	pooled

Common characteristics:

- Ionic:** High melting points, most dissolve in water, conduct electricity when dissolved in water, brittle.
- Covalent:** Low melting points, most do not dissolve in water, do not conduct electricity when dissolved in water.
- Polar covalent:** Medium melting points, some dissolve in water, do not conduct electricity when dissolved in water.
- Metallic:** Soft, conduct heat and electricity, do not dissolve in water.

Bond Polarity

When non-metals bond covalently with a large difference in electronegativity.

- Absolute value of differences:
 - 0 - 0.4 = covalent
 - 0.5 - 1.4 = polar covalent
 - 1.5 - 4 = ionic

Symbolized with an arrow pointing towards the more electronegative element and a crossed tail by the less electronegative element.

Example:
 $\text{C} \rightarrow \text{N}$ Electronegativity difference of 0.5 = polar bond



Isomers and Resonance

Isomer examples:



Resonance examples:



Sigma and Pi bonds

How many sigma and pi bonds are in the following?



6 sigma bonds & 3 pi bonds

Hybridization

# of sigma bonds	Hybridization	Formed from
2	sp	1 s and 1 p
3	sp^2	1 s and 3 p
4	sp^3	1 s and 3 p

How to Use This Cheat Sheet: These are the keys related this topic. Try to read through it carefully twice then rewrite it out on a blank sheet of paper. Review it again before the exams.

Chemistry The Physical Setting 2013 Answers

**Guangzhi Sun,Xiaofei Yu,Yuanchun
Zou,Guobin Fu**



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Sustainable Carbon Capture Humbul Suleman, Philip Loldrup Fosbøl, Rizwan Nasir, Mariam Ameen, 2022-02-16 A comprehensive resource on different aspects of sustainable carbon capture technologies including recent process developments environmentally friendly methods and roadmaps for implementations It discusses also the socio economic and policy aspects of carbon capture and the challenges opportunities and incentives for change with a focus on industry policy and governmental sector Through applications in various fields of environmental health and four selected case studies from four different practical regimes of carbon capture the book provides guidelines for sustainable and responsible carbon capture and addresses current and future global energy environment and climate concerns

Membrane Proteins in Aqueous Solutions Jean-Luc Popot, 2018-06-08 This book is the first to be entirely devoted to the challenging art of handling membrane proteins out of their natural environment a key process in biological and pharmaceutical research but one plagued with difficulties and pitfalls Written by one of the foremost experts in the field *Membrane Proteins in Aqueous Solutions* is accessible to any member of a membrane biology laboratory After presenting the structure functions dynamics synthesis natural environment and lipid interactions of membrane proteins the author discusses the principles of extracting them with detergents the mechanisms of detergent induced destabilization countermeasures and recent progress in developing detergents with weaker denaturing properties Non conventional alternatives to detergents including bicelles nanodiscs amphipathic peptides fluorinated surfactants and amphipols are described and their relative advantages and drawbacks are compared The synthesis and solution properties of the various types of amphipols are presented as well as the formation and properties of membrane protein amphipol complexes and the transfer of amphipol trapped proteins to detergents nanodiscs lipidic mesophases or living cells The final chapters of the book deal with applications membrane protein in vitro folding and cell free expression solution studies NMR crystallography electron microscopy mass spectrometry amphipol mediated immobilization of membrane proteins and biomedical applications Important features of the book include introductory sections describing foundations as well as the state of the art for each of the biophysical techniques discussed and topical tables which organize a widely dispersed literature Boxes and annexes throughout the book explain technical aspects and twelve detailed experimental protocols ranging from in vitro folding of membrane proteins to single particle electron cryomicroscopy have been contributed by and commented on by experienced users *Membrane Proteins in Aqueous Solutions* offers a concise accessible introduction to membrane protein biochemistry and biophysics as well as comprehensive coverage of the properties and uses of conventional and non conventional surfactants It will be useful both in basic and applied research laboratories and as a teaching aid for students instructors researchers and professionals within the field

Sustainable Groundwater and Environment: Challenges and Solutions Peiyue Li, Xiaodong He, Jianhua Wu, Vetrimurugan Elumalai, 2025-02-17 This book addresses this challenge head on focusing on the intersection of

groundwater sustainability and environmental health It explores the complex relationship between groundwater management and environmental integrity delving into topics like the impacts of groundwater extraction on ecosystems the role of groundwater in climate change adaptation and mitigation and the effects of surface water groundwater exchange on groundwater quality The book also highlights the latest scientific research technological advancements in the field of sustainable groundwater management Moreover this book not only identifies the challenges but also proposes viable solutions It presents a range of practices innovative solutions and policy recommendations to manage groundwater resources effectively while minimizing environmental impact These will be drawn from a variety of contexts and case studies making the book a comprehensive guide for researchers policy makers practitioners and students in the field By providing a holistic perspective on sustainable groundwater management this book contributes significantly to our collective efforts toward achieving water security and environmental sustainability Wetlands and Water Resources: Ecological Solutions for the Environment

Guangzhi Sun,Xiaofei Yu,Yuanchun Zou,Guobin Fu,2023-04-12 Wetlands are usually located in complex landscapes Multiple ecosystem services delivered by wetlands to human well being should be considered in broader contexts of wetland functions water quality water resource management and wetland conservation Natural wetlands are kidneys and sponges of the earth that remove water pollutants and provide hydrological regulation functions vital to the environment Where the functions of the wetlands are weakened actions of rehabilitations are needed to revitalize those functions While worldwide many natural wetlands have disappeared due to climate change and or human activities constructed wetlands which may be seen as artificial kidneys sponges of the earth have been increasingly built to replicate the functions of the natural wetlands This Research Topic aims to collect and present the latest research developments in natural and constructed wetlands in relation to either water resource functions or water quality Articles published in this Research Topic will showcase new understandings of water balances and flow patterns vegetation characteristics and pollutant transformations in wetland systems innovative designs and efficiencies of wastewater treatment wetlands and technologies that revitalize the ecosystem functions of natural wetlands These articles are expected to bring benefits to research communities water and wastewater industries and government agencies that deal with wetland water management and conservation High quality Original Research and Review articles in this field are all welcome for submission to this Research Topic Research interests include but are not limited to the following areas 1 Water balances and flow patterns in natural wetland systems 2 The hydrological regulation functions and importance of natural wetlands in their environmental context 3 Vegetation characteristics of natural wetlands and their functional importance in wetland water pollution control 4 Techniques for the rehabilitation of degraded wetlands 5 The role of constructed wetlands in wastewater treatment 6 Innovative designs and efficiencies of wastewater treatment wetlands 7 The modeling of hydrological and or pollutant dynamics in wetland systems 8 The ecohydrology of wetlands 9 The role of wetlands in the global carbon cycle 10 Impacts of

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