

Chemistry – Unit 6 Objectives Describing Substances

1. Classify pure substances and mixtures and molecular and ionic compounds. Describe the structure of pure substances and ionic compounds. a. Draw space-filling models for: i. pure substances with a crystal lattice structure ii. simple molecules iii. ionic compounds showing the pattern of alternating ionic charges	a. pure substances b. ionic compounds c. simple ion size d. simple ion charge e. ion size f. ion charge	<table><tr><td>pure substances</td><td>10</td><td>10</td><td>20</td></tr><tr><td>ionic compounds</td><td>10</td><td>10</td><td>20</td></tr><tr><td>simple ion size</td><td>10</td><td>10</td><td>20</td></tr><tr><td>simple ion charge</td><td>10</td><td>10</td><td>20</td></tr><tr><td>ion size</td><td>10</td><td>10</td><td>20</td></tr><tr><td>ion charge</td><td>10</td><td>10</td><td>20</td></tr></table>	pure substances	10	10	20	ionic compounds	10	10	20	simple ion size	10	10	20	simple ion charge	10	10	20	ion size	10	10	20	ion charge	10	10	20																
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2. Identify physical states of compounds and mixtures. a. List 5 substances that represent the states	a. b. c. d.																																									
3. Compare the structure, physical and chemical properties and physical and chemical changes. a. Clearly state at least 3 following as physical or chemical changes b. Clearly state at least 3 following as a physical or chemical property	a. Describing states of matter Describing matter Describing chemical change Describing physical change b. matter c. density Describing matter Describing physical change Describing chemical change Describing physical property	<table><tr><td>Describing states of matter</td><td>10</td><td>10</td><td>20</td></tr><tr><td>Describing matter</td><td>10</td><td>10</td><td>20</td></tr><tr><td>Describing chemical change</td><td>10</td><td>10</td><td>20</td></tr><tr><td>Describing physical change</td><td>10</td><td>10</td><td>20</td></tr></table> <table><tr><td>matter</td><td>10</td><td>10</td><td>20</td></tr><tr><td>density</td><td>10</td><td>10</td><td>20</td></tr><tr><td>Describing matter</td><td>10</td><td>10</td><td>20</td></tr><tr><td>Describing physical change</td><td>10</td><td>10</td><td>20</td></tr><tr><td>Describing chemical change</td><td>10</td><td>10</td><td>20</td></tr><tr><td>Describing physical property</td><td>10</td><td>10</td><td>20</td></tr></table>	Describing states of matter	10	10	20	Describing matter	10	10	20	Describing chemical change	10	10	20	Describing physical change	10	10	20	matter	10	10	20	density	10	10	20	Describing matter	10	10	20	Describing physical change	10	10	20	Describing chemical change	10	10	20	Describing physical property	10	10	20
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Describing chemical change	10	10	20																																							
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4. Compare pure substances and mixtures and molecular and ionic compounds. Describe the structure of pure substances and ionic compounds. a. Draw 5 sets of chemical structures b. Explain how the law applies to	a. b.																																									

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**John Robertson Dunlap, Henry Clemens
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