

Chemistry Unit 3

Study Guide

Chemical Nomenclature

Chemical formula

- a. a way of presenting information about the chemical proportions of atoms that constitute a particular chemical compound or molecule, using chemical element symbols, numbers, and sometimes also other symbols

Subscript

- a. A subscript or superscript is a character that is set slightly below or above the normal line of type, respectively. It is usually smaller than the rest of the text. Subscripts appear at or below the baseline, while superscripts are above.

Ion

- a. Ion, any atom or group of atoms that bears one or more positive or negative electrical charges.

Anion

- a. a negatively charged ion

Cation

- a. A positively charged ion

Polyatomic ion

- a. The prefix poly- means many, so a polyatomic ion is an ion that contains more than one atom.

Understand the use of Roman Numerals in ionic compounds

- a. Roman numerals are used in naming ionic compounds when the metal cation forms more than one ion.

List the diatomic molecules

- a. Hydrogen (H_2)
- b. Nitrogen (N_2)
- c. Oxygen (O_2)
- d. Fluorine (F_2)
- e. Chlorine (Cl_2)
- f. Iodine (I_2)
- g. Bromine (Br_2)

Be able to name and write chemical formulas for ionic, covalent, and acidic compounds

- Properly use prefixes
- Properly use Roman Numerals
- Properly use endings (-ide, -ate, -ite, -ous, -ic)

Determine the type of compound made from a pair of elements or polyatomic ions

Identify the ionic charge of transition metals in compounds

- a. Because most transition metals have two valence electrons, the charge of $2+$ is a very common one for their ions.

Identify compounds as ionic, covalent, or acidic

- a. Compounds containing two elements (so called binary compounds) can either have ionic or covalent bonding. If a compound is made from a metal and a nonmetal, its bonding will be ionic. If a compound is made from two non-metals, its bonding will be covalent.

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