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Nuclear Decay

Types of Radiation

Complete the table below to compare the properties of alpha, beta, positron, and gamma radiation.

Name	Alpha	Beta	Positron	Gamma
Greek letter	α	β-	β+	Y
Symbol	*He	e 1e	e	N/A
Composition	2p°, 2n°	electron	anti-electron	EM wave
Charge	2+	1-	1+	0
Stopped by	paper, skin, clothing	wood, glass	wood, glass	thick concrete or lead

Nuclear Decay Reactions

In problems 2 - 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.

$$B_2 = {}^{14}C \rightarrow {}^{14}N + {}^{6}e$$

$$\underline{D}$$
 3. $^{32}_{17}CI \rightarrow ^{0}_{+1}e + ^{32}_{16}S$

$$A$$
 4. $^{218}_{54}Po \rightarrow ^{4}_{2}He + ^{214}_{82}Pb$

$$C$$
 5. $^{142}_{61}Pm + ^{0}_{-1}e \rightarrow ^{142}_{60}Nd$

A. alpha emission

B. beta emission

C. electron capture

D. positron emission

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

- Decay of radium-223 by alpha (α) emission.
- Decay of iodine-126 by positron (β+) emission.

$$^{223}_{88}Ra \rightarrow ^{4}_{2}He + ^{219}_{86}Rn$$

$$^{126}_{53}I \rightarrow ^{0}_{-1}e + ^{126}_{52}Te$$

8. How are the atomic number and mass number of a nuclide affected by positron emission?

Atomic number decreases by one. Mass number stays the same.

- How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?
 Both atomic number and mass number stay the same since gamma radiation has no mass or charge.
- 10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why? The reactants have a slightly greater mass. In a nuclear reaction, a small amount of mass is converted to energy according to the equation E = mc². The difference in mass is referred to as the mass defect.

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