

Chemistry Study Guide

Periodic Table of the Elements

<div>1</div> <div>H</div>	<div>2</div> <div>He</div>																	<div>18</div> <div>Ar</div>	<div>19</div> <div>K</div>	<div>20</div> <div>Ca</div>											<div>36</div> <div>Kr</div>	<div>37</div> <div>Rb</div>	<div>38</div> <div>Sr</div>											<div>54</div> <div>Xe</div>	<div>55</div> <div>Cs</div>	<div>56</div> <div>Ba</div>															<div>86</div> <div>Rn</div>																															
<div>3</div> <div>Li</div>	<div>4</div> <div>Be</div>																	<div>12</div> <div>Mg</div>	<div>13</div> <div>Al</div>	<div>14</div> <div>Si</div>	<div>15</div> <div>P</div>	<div>16</div> <div>S</div>	<div>17</div> <div>Cl</div>	<div>18</div> <div>Ar</div>	<div>19</div> <div>K</div>	<div>20</div> <div>Ca</div>	<div>21</div> <div>Sc</div>	<div>22</div> <div>Ti</div>	<div>23</div> <div>V</div>	<div>24</div> <div>Cr</div>	<div>25</div> <div>Mn</div>	<div>26</div> <div>Fe</div>	<div>27</div> <div>Co</div>	<div>28</div> <div>Ni</div>	<div>29</div> <div>Cu</div>	<div>30</div> <div>Zn</div>	<div>31</div> <div>Ga</div>	<div>32</div> <div>Ge</div>	<div>33</div> <div>As</div>	<div>34</div> <div>Se</div>	<div>35</div> <div>Br</div>	<div>36</div> <div>Kr</div>	<div>37</div> <div>Rb</div>	<div>38</div> <div>Sr</div>	<div>39</div> <div>Y</div>	<div>40</div> <div>Zr</div>	<div>41</div> <div>Nb</div>	<div>42</div> <div>Mo</div>	<div>43</div> <div>Tc</div>	<div>44</div> <div>Ru</div>	<div>45</div> <div>Rh</div>	<div>46</div> <div>Pd</div>	<div>47</div> <div>Ag</div>	<div>48</div> <div>Cd</div>	<div>49</div> <div>In</div>	<div>50</div> <div>Sn</div>	<div>51</div> <div>Sb</div>	<div>52</div> <div>Te</div>	<div>53</div> <div>I</div>	<div>54</div> <div>Xe</div>	<div>55</div> <div>Cs</div>	<div>56</div> <div>Ba</div>	<div>57</div> <div>La</div>	<div>58</div> <div>Ce</div>	<div>59</div> <div>Pr</div>	<div>60</div> <div>Nd</div>	<div>61</div> <div>Pm</div>	<div>62</div> <div>Sm</div>	<div>63</div> <div>Eu</div>	<div>64</div> <div>Gd</div>	<div>65</div> <div>Tb</div>	<div>66</div> <div>Dy</div>	<div>67</div> <div>Ho</div>	<div>68</div> <div>Er</div>	<div>69</div> <div>Tm</div>	<div>70</div> <div>Yb</div>	<div>71</div> <div>Lu</div>	<div>72</div> <div>Hf</div>	<div>73</div> <div>Ta</div>	<div>74</div> <div>W</div>	<div>75</div> <div>Re</div>	<div>76</div> <div>Os</div>	<div>77</div> <div>Ir</div>	<div>78</div> <div>Pt</div>	<div>79</div> <div>Au</div>	<div>80</div> <div>Hg</div>	<div>81</div> <div>Tl</div>	<div>82</div> <div>Pb</div>	<div>83</div> <div>Bi</div>	<div>84</div> <div>Po</div>	<div>85</div> <div>At</div>	<div>86</div> <div>Rn</div>
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🎯 Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🧪

💡 Matter:

Anything that has **mass** and **occupies space** 🧪🌍

Examples: air, gases, liquids, solids!

🔬 Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

📌 Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

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Table of Contents Chemistry Study Guide For Oklahoma

1. Understanding the eBook Chemistry Study Guide For Oklahoma
 - The Rise of Digital Reading Chemistry Study Guide For Oklahoma
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide For Oklahoma
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide For Oklahoma
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide For Oklahoma

- Personalized Recommendations
 - Chemistry Study Guide For Oklahoma User Reviews and Ratings
 - Chemistry Study Guide For Oklahoma and Bestseller Lists
5. Accessing Chemistry Study Guide For Oklahoma Free and Paid eBooks
 - Chemistry Study Guide For Oklahoma Public Domain eBooks
 - Chemistry Study Guide For Oklahoma eBook Subscription Services
 - Chemistry Study Guide For Oklahoma Budget-Friendly Options
 6. Navigating Chemistry Study Guide For Oklahoma eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Study Guide For Oklahoma Compatibility with Devices
 - Chemistry Study Guide For Oklahoma Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide For Oklahoma
 - Highlighting and Note-Taking Chemistry Study Guide For Oklahoma
 - Interactive Elements Chemistry Study Guide For Oklahoma
 8. Staying Engaged with Chemistry Study Guide For Oklahoma
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide For Oklahoma
 9. Balancing eBooks and Physical Books Chemistry Study Guide For Oklahoma
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide For Oklahoma
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Study Guide For Oklahoma
 - Setting Reading Goals Chemistry Study Guide For Oklahoma
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Study Guide For Oklahoma

- Fact-Checking eBook Content of Chemistry Study Guide For Oklahoma
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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