

Summer Assignment Chapter 2

PROBLEMS pg. 69 #7, 11, 13, 15, 17, 20, 21, 22, 27, 28, 30, 31, 33, 35, 38, 40, 41, 43, 45, 47, 49, 55, 56, 57, 58, 60, 61, 64, 67, 69, 71, 87

- 7 The diameter of a neutral helium atom = 1×10^2 pm. Suppose that we could line up helium atoms side by side in contact with one another. Approximately how many atoms would it take to make the distance from end to end 1 cm?

First, convert 1 cm to picometers. $1 \text{ cm} \times \frac{0.01 \text{ m}}{1 \text{ cm}} \times \frac{1 \text{ pm}}{1 \times 10^{-12} \text{ m}} = 1 \times 10^{10} \text{ pm}$

$$? \text{ He atoms} = (1 \times 10^{10} \text{ pm}) \times \frac{1 \text{ He atom}}{1 \times 10^2 \text{ pm}} = 1 \times 10^8 \text{ He atoms}$$

- 11 What do we call atoms of the same elements with different mass numbers? Isotopes.

- 13 What is the mass number of an iron atom that has 28 neutrons?

For iron, the atomic number Z is 26. Therefore the mass number A is: $A = 26 + 28 = 54$

- 15 For the following atoms, determine the number of protons and neutrons.

Isotope	${}^3_2\text{He}$	${}^4_2\text{He}$	${}^{24}_{12}\text{Mg}$	${}^{25}_{12}\text{Mg}$	${}^{48}_{22}\text{Ti}$	${}^{79}_{35}\text{Br}$	${}^{195}_{78}\text{Pt}$
No. Protons	2	2	12	12	22	35	78
No. Neutrons	1	2	12	13	26	44	117

- 17 Give the symbol for the following isotopes:

(a) $Z = 11$ $A = 23$ ${}^{23}_{11}\text{Na}$ (b) $Z = 28$ $A = 64$ ${}^{64}_{28}\text{Ni}$

- 20 State two differences between a metal and a nonmetal.

Metals conduct heat and electricity, are shiny & malleable; nonmetals are poor conductors, generally dull and brittle.

- 21 Write the names and symbols for four elements in each of the following categories:

(a) nonmetal: phosphorous P, bromine Br, neon, Ne, carbon C

(b) metal: iron Fe, cesium Cs, iridium, Ir, praseodymium Pr

(c) metalloid: silicon Si, germanium Ge, boron B, antimony Sb

- 22 Define, with two examples, the following terms:

(a) Alkali metals: (1st column) Sodium, rubidium;

(b) alkaline earth metals: (2nd column) barium, beryllium;

(c) halogens: (7th column) fluorine, iodine;

(d) noble gases: (8th column) xenon, krypton

- 27 What is the difference between an atom and a molecule?

An atom is the smallest representative unit of an element.

A molecule is the smallest representative unit of a (non-ionic) compound, composed of two or more atoms chemically bonded.

- 28 What are allotropes? Allotropes are different physical forms of the same element.

Examples? graphite and diamond are two allotropes of carbon.

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Daniela Niemeyer



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