

Chemistry: The Central Science, 14e (Brown, et al.)
Chapter 2 Atoms, Molecules, and Ions

2.1 Multiple-Choice Questions

- 1) A molecule of water contains hydrogen and oxygen in a 1:8 ratio by mass. This is a statement of _____.
- A) the law of multiple proportions
 - B) the law of constant composition
 - C) the law of conservation of mass
 - D) the law of conservation of energy
 - E) none of the above

Answer: B

Diff: 2 Var: 1 Page Ref: Sec. 2.1

LO: 2.1

GO: G2

- 2) Which one of the following is not one of the postulates of Dalton's atomic theory?
- A) Atoms are composed of protons, neutrons, and electrons.
 - B) All atoms of a given element are identical; the atoms of different elements are different and have different properties.
 - C) Atoms of an element are not changed into different types of atoms by chemical reactions; atoms are neither created nor destroyed in chemical reactions.
 - D) Compounds are formed when atoms of more than one element combine; a given compound always has the same relative number and kind of atoms.
 - E) Each element is composed of extremely small particles called atoms.

Answer: A

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- 3) Consider the following selected postulates of Dalton's atomic theory:
- (i) Each element is composed of extremely small particles called atoms.
 - (ii) Atoms are indivisible.
 - (iii) Atoms of a given element are identical.
 - (iv) Atoms of different elements are different and have different properties.

Which of the postulates is(are) no longer considered valid?

- A) (i) and (ii)
- B) (ii) only
- C) (ii) and (iii)
- D) (iii) only
- E) (iii) and (iv)

Answer: C

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Chemistry The Central Science Chapter Review Questions

Thomas Gardner



Chemistry The Central Science Chapter Review Questions:

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